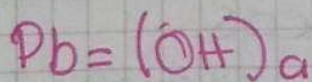
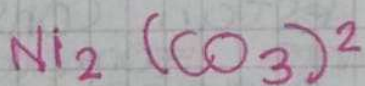


## Solution.

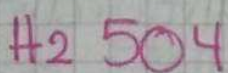
$$\begin{aligned}H &= 1 \cdot 3 = 3 \div 98 = 0.030 \cdot 100 = 3.0\% H \\P &= 31 \cdot 1 = 31 \div 98 = 0.316 \cdot 100 = 31.6\% P \\O &= 16 \cdot 4 = \frac{64}{98 \text{ g/mol}} \div 98 = 0.653 \cdot 100 = 65.3\% O\end{aligned}$$



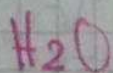
$$\begin{aligned}Pb &= 207 \cdot 1 = 207 \div 275 = 0.152 \cdot 100 = 15.2\% Pb \\O &= 16 \cdot 4 = 64 \div 275 = 0.232 \cdot 100 = 23.2\% O \\H &= 1 \cdot 4 = \frac{4}{27 \text{ g/mol}} \div 275 = 0.014 \cdot 100 = 1.4\% H\end{aligned}$$



$$\begin{aligned}Ni &= 59 \cdot 2 = 118 \div 111 = 0.531 \cdot 100 = 53.1\% Ni \\C &= 12 \cdot 3 = 36 \div 111 = 0.324 \cdot 100 = 32.4\% C \\O &= 16 \cdot 3 = 48 \div 111 = 0.144 \cdot 100 = 14.4\% O\end{aligned}$$



$$\begin{aligned}H &= 1 \cdot 2 = 2 \div 98 = 0.020 \cdot 100 = 2.0\% H \\S &= 32 \cdot 1 = 32 \div 98 = 0.326 \cdot 100 = 32.6\% S \\O &= 16 \cdot 4 = \frac{64}{98 \text{ g/mol}} \div 98 = 0.653 \cdot 100 = 65.3\% O\end{aligned}$$



$$\begin{aligned}H &= 1 \cdot 2 = 2 \div 18 = 0.111 \cdot 100 = 11.1\% H \\O &= 16 \cdot 1 = \frac{16}{18 \text{ g/mol}} \div 18 = 0.888 \cdot 100 = 88.8\% O\end{aligned}$$