Solucion.

H = 1 • 3 = 3 ÷ 98 = 0.030 · 100 = 3.0% H

P = 31 • 1 = 31 ÷ 98 = 0.316 · 100 = 31.6% P

O = 16 · 4 = 64 / 98 g/mol ÷ 98 = 0.653 · 100 = 65.3% O

$Pb(OH)_a$

Pb = 207 • 1 = 207 ÷ 275 = 0.152 · 100 = 15.2 Pb

O = 16 · 4 : 64 ÷ 275 = 0232 · 100 = 2320

H = 1 · 4 : 4 / 27 g/mol ÷ 275 = 0.014 · 100 = 1.4 H

$Ni_2(CO_3)^2$

Ni = 59 · 1 = 59 ÷ 111 = 0,531 · 100 = 5.31 Ni

C = 12 · 3 = 36 ÷ 111 = 0324 · 100 = 3.24 C

O = 16 3 · 3 = 16 / 114 g/mol ÷ 111 = 0.144 · 100 = 144.0

H_2SO_4

H: 1 · 2 = 2 ÷ 98 = 0.020 • 100 = 0.20 H

S · 32 · 1 = 32 ÷ 98 = 0.326 • 100 = 3.26 S

O · 16 · 4 = 64 / 98 g/mol ÷ 98 = 0.653 • 100 = 6.530

H_2O

H = 1 · 2 = 2 ÷ 18 = 0.111 · 100 = 1.11

O = 16 · 1 = 16 / 18 g/mol ÷ 18 = 0.888 · 100 = 888