

ALZATI DMS

C = 40%
H = 6,7%
O = 53,3%

masa = 100g

$$C = 40 \div 12 \text{ g/mol} = 3,3 \text{ mol} \div 3,3 = 1$$

$$H = 6,7 \div 1 \text{ g/mol} = 6,7 \text{ mol} \div 3,3 = 2$$

$$O = 53,3 \div 16 \text{ g/mol} = 3,3 \text{ mol} \div 3,3 = 1$$

Formula empirica = CH_2O

$$C = 12 \times 1 = 12$$

$$H = 1 \times 2 = 2$$

$$O = 16 \times 1 = \frac{16}{30}$$

$$n = 90 \text{ g} / 30 = 3$$

$$3 \times 1 = 3$$

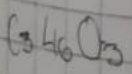
$$3 \times 2 = 6$$

$$3 \times 2 = 6$$

$$3 \times 2 = 6$$

$$3 \times 1 = 3$$

Formula molecular



2) C = 92,8%
 H = 6,3%
 Cl = 55,8%
 masa = 127g

$$C = 92,8 \div 12 \text{ g/mol} = 7,73 \text{ mol} \div 7,73 = 1$$

$$H = 6,3 \div 1 \text{ g/mol} = 6,3 \text{ mol} \div 7,73 = 0,81$$

$$Cl = 55,8 \div 35 \text{ g/mol} = 1,59 \text{ mol} \div 7,73 = 0,20$$

Formula empirica = $(C_1 H_0,81 Cl_0,20)$

$$C = 12 \times 1 = 12$$

$$H = 1 \times 0,81 = 0,81$$

$$Cl = 35 \times 0,20 = 7$$

$$n = 127 \text{ g} / 63 = 2$$

$$2 \times 12 = 24$$

$$2 \times 0,81 = 1,62$$

$$2 \times 7 = 14$$

Formula molecular

