

SOLUTION

$$H: 1 \cdot 3 = 3 \div 98 = 0.030 \cdot 100 = 3.0\% H$$

$$P: 31 \cdot 1 = 31 \div 98 = 0.316 \cdot 100 = 31.6\% P$$

$$O: 16 \cdot 4 = \frac{64}{98 \text{ g/mol}} \div 98 = 0.653 \cdot 100 = 65.3\% O$$

$$\bullet Pb = 207 \cdot 1 = 207 \div 275 = 0.152 \cdot 100 = 15.2\% Pb$$

$$O = 16 \cdot 4 = 64 \div 275 = 0.232 \cdot 100 = 25.2\% O$$

$$H = \frac{1 \cdot 4}{275 \text{ g/mol}} \div 275 = 0.014 \cdot 100 = 1.4\% H$$

$$\bullet Ni_2(CO_3)_3$$

$$Ni = 59 \cdot 2 = 118 \div 114 = 0.531 \cdot 100 = 53.1\% Ni$$

$$C = 12 \cdot 3 = 36 \div 114 = 0.324 \cdot 100 = 32.4\% C$$

$$O = 16 \cdot 3 = \frac{48}{114} \div 114 = 0.144 \cdot 100 = 14.4\% O$$

H_2SO_4

$$H = 1 \cdot 2 = 2 = 98 = 0.020 \cdot 100 = 0.20 \quad H$$

$$S = 32 \cdot 1 = 32 = 98 = 0.326 \cdot 100 = 3.265$$

$$O = 16 \cdot 4 = \frac{64}{98 \text{ g/mol}} = 0.653 \cdot 100 = 6.530$$

H_2O

$$H = 1 \cdot 2 = 2 = 18 = 0.111 \cdot 100 = 1.11$$

$$O = 16 \cdot 1 = \frac{16}{18 \text{ g/mol}} = 0.888 \cdot 100 = 8.88$$

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