$$\frac{P_1 V}{T_1} = \frac{P_2 V_2}{T_2}$$

$V_1 = 440 \, ml = 0,44 L$  $\qquad$  $V_2 = 5,6 L$

$T_1 = 70°C = 343,15°K$  $\qquad$  $T_2 = 100°C = 373,15°K$

$P_1 = 920 \, mm \, Hg$  $\qquad$  $P_2 = ?$

$$P_2 = \frac{P_1 V_1 T_2}{T_1 V_2}$$

$$P_2 = \frac{920 \, mmHg}{343,15 \, K} \cdot \frac{0,44 L}{5,6 L} \cdot 373,15 \, K$$

$P_2 = 78,60531629 \, mm \, Hg$  o  $0,1034280478 \, atm$

① $V_1 \times P_1 = V_2 \times P_2$

$\phantom{①} T_1 \phantom{\times P_1 = V_2} T_2$

Dato:

$V_1 = 205/1000 = 0.205 L$

$T_1 = 20 + 273 = 293 K$

$P_1 = 1.5 atm$

$V_2 = ?$

$T_2 = 60 + 273 = 333 K$

Res $\phantom{.} P_2 \times T_1$

$V_2 = 0.205 \times 1.0 satm \times 337 K$

$\phantom{V_2 =} 2.4 atm \times 293 k$

$V_2 = 0.106$ Litros de gas.

③ $H = 2.4 \times 10 - 6m$

$PH = 5.619$

$POH = -24 + 74$

$Poh = 8.381$

$Oh = 4.168 \times 10^{-4}$

④ $cH = 5.357$

$Ph = -5.357 + 74$

$Ph = 8.649m$

$H = 2.24 \times 10^{-4}$

$P^2ch\ 5.357$

21- Sep -20..

Comproda las caracteristicas y los procedimientos
de TH

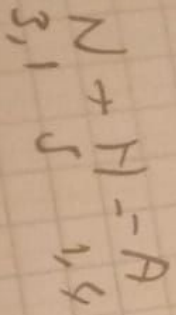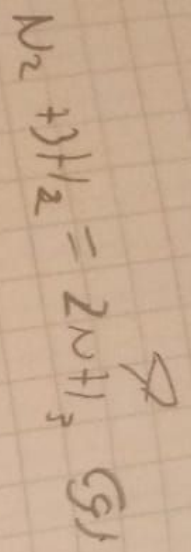calcular POH y la concentración de la solución si el PH
es de 7.2

PH = 7,2

POH = 14 - PH

= 6 8

= SHIft + log ÷ PoH - PH

= 1.68

$$N_2 + 3H_2 = 2NH_3 \text{ (g)}$$

$$\boxed{\frac{(NH_3)^2}{N_2 \times (H_2)^3}}$$

$$N + H = A$$
$$3,1 \quad 5 \quad 1,4$$

$$\frac{A}{N+H} \rightarrow \frac{1,4^2}{3,1 \times 5^3} = \frac{1,96}{5 \times 125} = \frac{1,96}{} = 0,005$$