FORMULA EMPIRICA.

1. $\text{mol C} = \frac{92.3\text{g}}{12\text{g/mol}} = 7.691\text{ mol}$ $7.691/7.7 = 1.$

$\text{mol H} = \frac{7,7\text{g}}{1\text{g/mol}} = 7.7\text{ mol}$ $7.7/7.7 = 1.$

- CH

2. $\text{mol Na} = \frac{32,4\text{g}}{23\text{g/mol}} = 1,408\text{ mol}$ $1,408/0.703 = 2$

$\text{mol S} = \frac{22,5\text{g}}{32\text{g/mol}} = 0.703\text{ mol}$ $0.703/0.703 = 1.$

$\text{mol O} = \frac{45,1\text{g}}{16\text{g/mo}} = 2,818\text{ mol}$ $2,818/0.703 = 4$

- Na_2SO_4

Scribe

3. $\text{mol C} = \frac{48\ g}{12\ g/mol} = 4$ mol 4/0.4 = 1

$\text{mol H} = \frac{4\ g}{1\ g/mol} = 4$ mol 4/0.4 = 1

$\text{mol N} = \frac{22,4\ g}{14\ g/mol} = 1,6$ mol 1.6/04 = 4

$\text{mol S} = \frac{12,8\ g}{32\ g\ mol} = 0,4$ mol 0,4/0,4 = 1.

4. $\text{mol N} = \frac{0,079\ g}{14\ g\ mol} = 5.642$ mol 5.642/001

$\text{mol O} = \frac{0,18\ 1g}{16\ g\ mol} = 0.011$ mol 0,011/0.011 =

• N_2O.

5. $\text{mol Na} = \frac{21,69}{23g/mol} = 0.939$ mol 0.934/0.939 = 1

mol Cl = 33,39 = 0,951 mol 0.951/0.939 = 1

$\text{mol O} = \frac{45,19}{16g/mo} = 2818$ mol 2.818/0.939 = 3

• $NaClO_3$.