

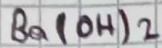
26

07

21

Lunes - Julio - 26 - 2021

1. Tenemos $\text{Ba}(\text{OH})_2$ con 210 gramos en 500 ml de solución. Calcula la Normalidad química.



$$\text{Ba} = 1 \times 137 = 137$$

$$m = 210 \text{ g}$$

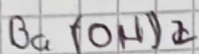
$$\text{O} = 2 \times 16 = 32$$

$$V = 500 \text{ ml}$$

$$\text{H} = 2 \times 1 = 2$$

196.5

24.5



$$N = \frac{(m)(\text{eq})}{(\text{PM})(V)}$$

$$m = 210 \text{ g}$$

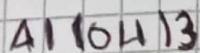
$$N = \frac{(210)(2)}{(171)(0.5)} = \frac{420}{85.5} = 4.91$$

$$\text{PM} = 171 \text{ g/mol}$$

$$V = 0.5 \text{ L}$$

Actividad

2. Obtén la Normalidad de $\text{Al}(\text{OH})_3$ con 450 ml y 9.50 gramos



$$\text{Al} = 1 \times 27 = 27$$

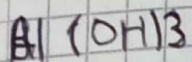
$$m = 9.50 \text{ g}$$

$$\text{O} = 3 \times 16 = 48$$

$$V = 450 \text{ ml}$$

$$\text{H} = 3 \times 1 = 3$$

78



$$N = \frac{(9.50)(3)}{(78)(0.45)} = \frac{28.5}{35.1} = 0.81$$

$$m = 9.50$$

$$\text{PM} = 78 \text{ g/mol}$$

$$V = 0.45$$

34.65

0.82