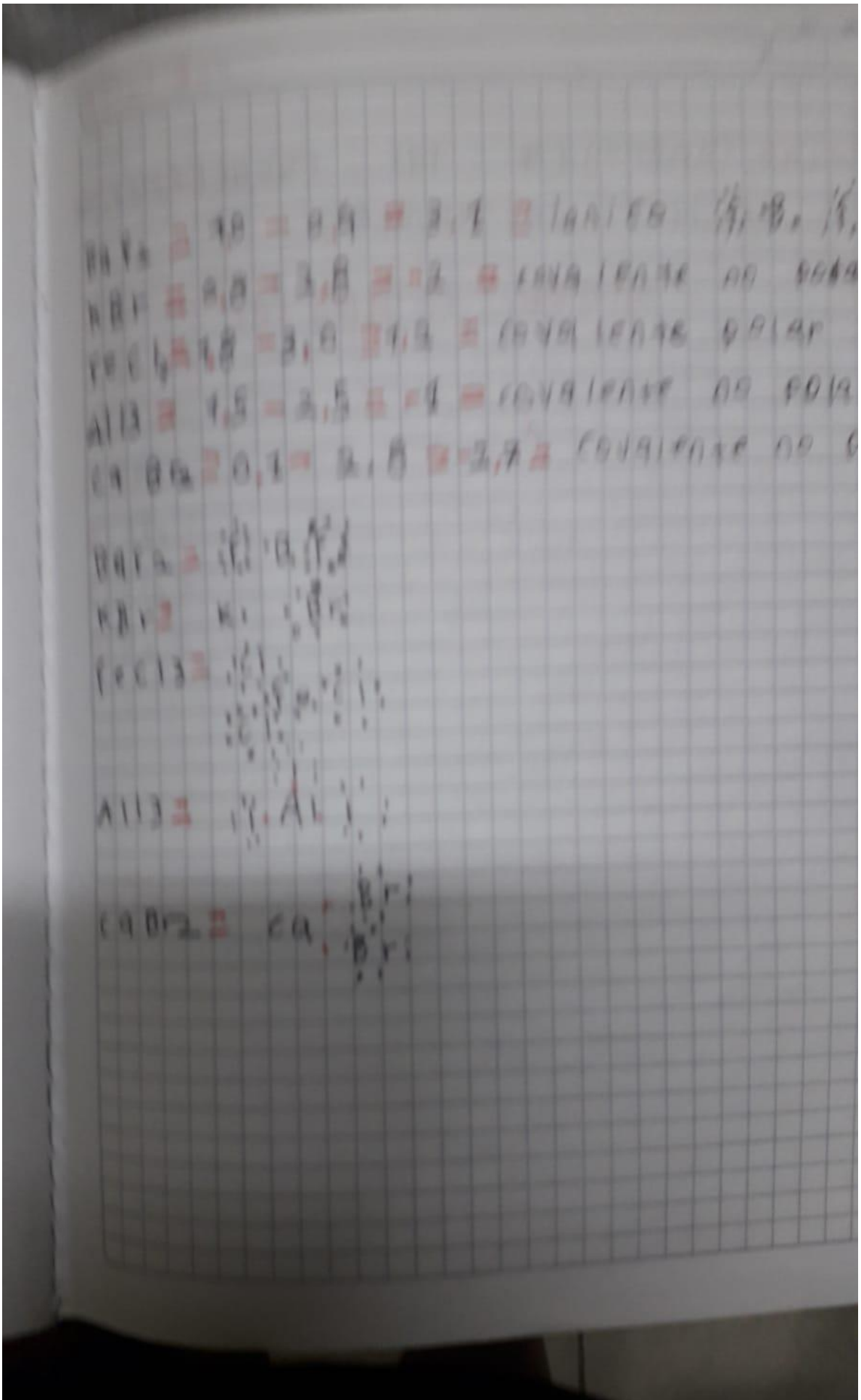
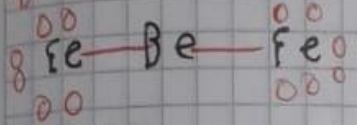


solucion

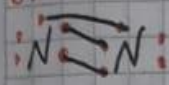
a)



3.98 - 1.57 = 2.41

Ionico

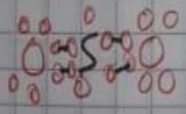
b)



covalente NO polar

c)

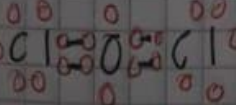
= 0,86



covalente polar

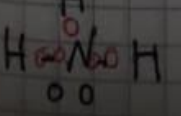
0,28

d)

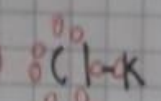


covalente no polar

e) 3,4 - 2,2 = 1,2



covalente Polar

f)  3,16 - 0,82 = 2,34

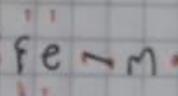
IONICO

g)  2,44

IONICO

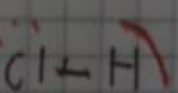
H)  0,4

COVALENTE NO POLAR

I)  0,52

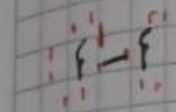
COVALENTE POLA

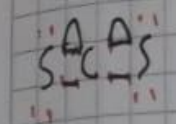
0,96

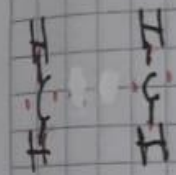
J) 

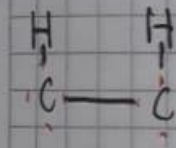
COVALENTE POLAR

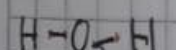
2)

 simple = F_2

 enlace doble = CS_2

 enlace simple = C_2H_4

 enlace simple = C_2H_2

 Enlace simple = H_2O

enlace simple = C_6H_6

enlace simple = NH_3