

Scribe

$$\textcircled{1} \text{ C} = 92,3 \%$$

$$\text{H} = 7,7 \%$$

$$\text{Mol C} = \frac{92,3 \text{ g}}{12 \text{ g/mol}} = 7,691 / 7,7 = 1$$

$$\text{Mol H} = \frac{7,7 \text{ g}}{1 \text{ g/mol}} = 7,7 / 7,7 = 1$$



$$\textcircled{2} \text{ Na} = 32,4 \%$$

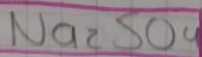
$$\text{S} = 22,5 \%$$

$$\text{O} = 45,1 \%$$

$$\text{Mol Na} = \frac{32,4 \text{ g}}{23 \text{ g/mol}} = 1,408 / 0,703 = 2$$

$$\text{Mol S} = \frac{22,5 \text{ g}}{32 \text{ g/mol}} = 0,703 / 0,703 = 1$$

$$\text{Mol O} = \frac{45,1 \text{ g}}{16 \text{ g/mol}} = 2,818 / 0,703 = 4$$



$$③ \text{ C} = 48 \%$$

$$\text{H} = 4 \%$$

$$\text{N} = 22,4 \%$$

$$\text{S} = 12,8 \%$$

$$\text{O} = 12,8 \%$$

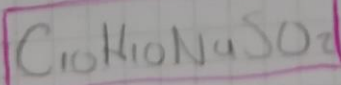
$$\text{Mol C} = \frac{48 \text{ g}}{12 \text{ g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol H} = \frac{4 \text{ g}}{14 \text{ g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol N} = \frac{22,4 \text{ g}}{14 \text{ g/mol}} = 1,6 / 0,4 = 4$$

$$\text{Mol S} = \frac{12,8 \text{ g}}{32 \text{ g/mol}} = 0,4 / 0,4 = 1$$

$$\text{Mol O} = \frac{12,8 \text{ g}}{16 \text{ g/mol}} = 0,8 / 0,4 = 2$$



$$④ \text{ N} = 0,079$$

$$\text{O} = 0,181$$

Scribe

$$\text{Mol N} = \frac{0,079 \text{ g}}{14 \text{ g/mol}} = 0,005 / 0,005 = 1$$

$$\text{Mol O} = \frac{0,181 \text{ g}}{16 \text{ g/mol}} = 0,011 / 0,005 = 2$$

**NO<sub>2</sub>**

③ Na = 21,6 %

Cl = 33,3 %

O = 45,1 %

$$\text{Mol Na} = \frac{21,6 \text{ g}}{23 \text{ g/mol}} = 0,939 / 0,939 = 1$$

$$\text{Mol Cl} = \frac{33,3 \text{ g}}{35 \text{ g/mol}} = 0,951 / 0,939 = 1$$

$$\text{Mol O} = \frac{45,1 \text{ g}}{16 \text{ g/mol}} = 2,818 / 0,939 = 3$$

**NaClO<sub>3</sub>**