

$$N = \frac{\text{equivalente gr soluto}}{\text{Litro de solución}}$$

$$N = \frac{(M) (ed)}{(PM) (V)}$$

$$1 \text{ Al(OH)}_3 \quad m = 9,50 \text{ g} \quad V = 450 \text{ ml} = 0,45 \text{ L}$$



$$PM: H: 3 \times 1 = 3$$

$$O: 3 \times 16 = 48$$

$$Al: 1 \times 27 = 27$$

$$PM: 78 \text{ g/mol}$$

~~N = ?~~

$$2 \quad m: 9,50 \text{ g}$$

$$PM: 78 \text{ g/mol}$$

$$V: 0,45 \text{ L}$$

$$N: ?$$

$$N = \frac{(9,50) (3)}{(78) (0,45)}$$

$$N = \frac{28,5}{35,1} = 0,81$$

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$$N = 0,811$$