

Solución.

$$N = \frac{\text{moles de soluto}}{\text{kg de solvente}}$$

$$M = \frac{\text{gramos}}{(\text{PM}) (\text{kg agua})}$$

$$\text{gramos} = (\text{m}) (\text{PM})$$

$$M = \frac{n (\text{soluto})}{\text{K agua}}$$

$$(\text{kg agua})$$

EJERCICIO 1

m:?

$$m = g / (\text{PM}) (\text{kg})$$

g: 95

$$m = 95 / (63) (0,025)$$

Agua: 25g 0,025 kg

$$m = 95 / 1,575$$

$\text{Pm} = \text{H} \quad 1 \times 1 = 1 = 1$

$$m = 60,3$$

$\text{N} \quad 1 \times 14 = 14 = 14$

$\text{O} \quad 3 \times 16 = 48 = 48$

$\underline{63 \text{ g/mol}}$

EJERCICIOS

M?

Pm

H 4 x 1

g. 15g

O 1 x 16

agua: 50g = 0,05kg

C 1 x 12

$m = g / (Pm) (kg)$

29 g/mol

$M = 15 / (29) / (0,05)$

$M = 15 / 1,45$

$M = 10,34$