

identificar comprender el procedimiento para obtener la constante de equilibrio.

Constante Equilibrio:

K: Constante equilibrio

$$K_c = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

A, B: productos. c, D: reactivos

[A]: Concentración de equilibrio de A en moles.

a: Numero de moles de A

L: liquido

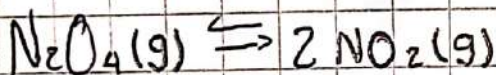
S: solido

AQ: acuoso

g: gaseoso

No

Si



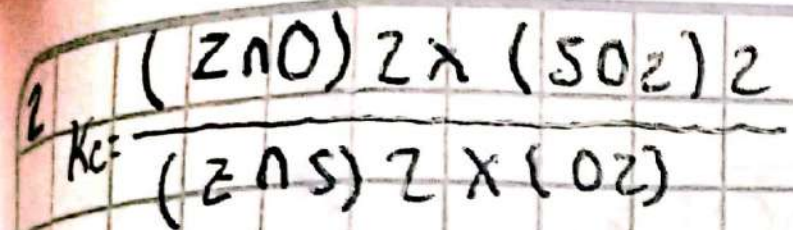
$$K_c = \frac{[NO_2]^2}{[N_2O_4]}$$

Ejercicio:

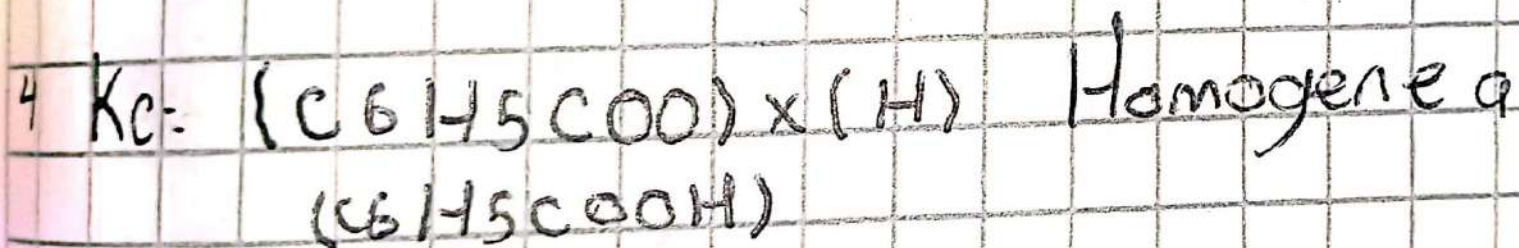
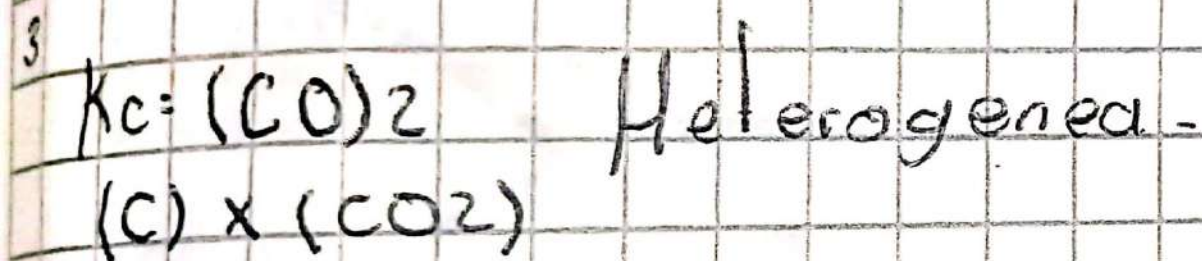
1.  $K_c = \frac{[NH_3]^x \times [H_2O]^4}{[NO_2]^2 \times [H_2]^7}$

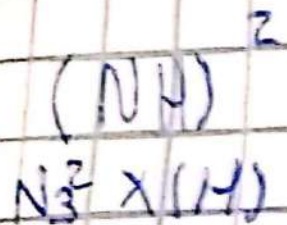
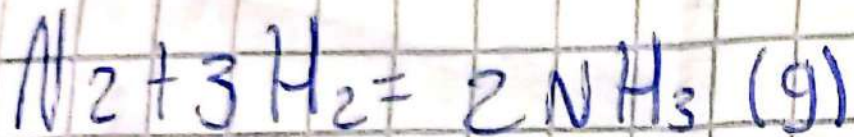
Heterogenea

2.  ~~$K_c = \frac{[2NO_2] \times [NO_2]^2}{[NO_2]^2}$~~



Heterogeneous.



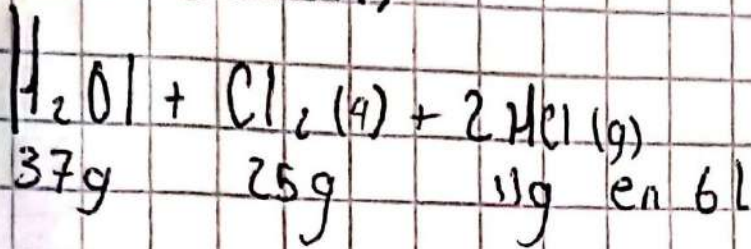


$$N + H = A$$
$$3.1 \quad 5 \quad 1.4$$

$$\frac{1.4^2}{3.1 \times 5^3} = \frac{196}{3.1 \times 125} = 0.005$$

Manuela

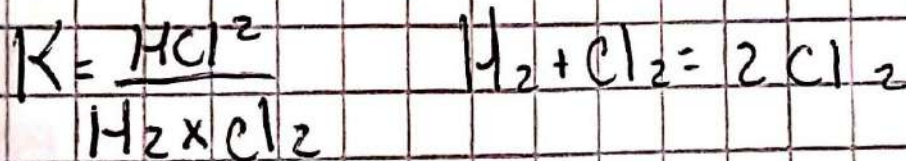
CSO



$$\text{H}_2: 1 \times 2 = 2 \rightarrow 37/2 \text{ g/mol} = 18,5 \rightarrow 18,5/6 = 3,08$$

$$\text{Cl}_2: 35 \times 2 = 70 \rightarrow 25/70 \text{ g/mol} = 0,357 \rightarrow 0,357/6 = 0,058$$

$$\text{HCl}: 36 \rightarrow 11/36 \text{ g/mol} = 0,30 \rightarrow 0,50/6 = 0,05$$



$\text{H}_2$	$\text{Cl}_2$	$\text{HCl}$	$K = \frac{(0,050)^2}{(3,08) \times (0,058)}$
37g	25g	11g	
3,08	0,058	0,05	$\frac{0,0025}{(3,08) \times (0,058)}$

$$K = 0,013$$

$$\frac{0,0025}{0,1789} = 0,013$$