Formula empírica

1 C = 92,3%

H = 7,7%

Mol C = $\frac{92,3\ g}{12\ g/mol}$ = 7,69 ≈ 7,69 = 1

Mol H = $\frac{7,7\ g}{1\ g/mol}$ = 7,7 ≈ 7,69 = 1

CH

2 Na = 32,4%

S = 22,5%

O = 45,1%

Mol Na = $\frac{32,4\ g}{22\ g/mol}$ = 1,47 / 1,40 = 1

Mol S = $\frac{22,5\ g}{16\ g/mol}$ = 1,40 / 1,40 = 1

Mol O = $\frac{45,1\ g}{16\ g/mol}$ = 2,81 / 1,40 = 2

$NaSO_2$

3 C = 48%

H = 4%

N = 22,4%

S = 12,8% Azufre

O = 12,8%

$C = \frac{48 g}{12 g/mol} = 4/0,4 = 10$

$H = \frac{4 g}{1 g/mol} = 4/0,4 = 10$

$N = \frac{22,4 g}{14 g/mol} = 1,6/0,4 = 4$

$S = \frac{12,8 g}{32 g/mol} = 0,4/0,4 = 1$

$O = \frac{12,8 g}{16 g/mol} = 0,8/0,4 = 2$

$C_{10}H_{10}N_4SO_2$

4 N = 0,079 g

O = 0,181 g

$N = \frac{0,079\ g}{14\ g/mol} = 0,001 / 0,00$

$O = \frac{0,181\ g}{16\ g/mol} = 0,01 / 0,00$

NO

5 Na = 21,6 %

Cl = 33,3 %

O = 45,1 %

$Na = \frac{21,6\ g}{22\ g/mol} = 0,98 / 0,95 = 1$

$Cl = \frac{33,3\ g}{35\ g/mol} = 0,95 / 0,95 = 1$

$O = \frac{45,1\ g}{16\ g/mol} = 2,81 / 0,95 = 2$

$NaClO_2$