

Formula empírica

1 C = 92,3%

H = 7,7%

$$\text{Mol } C = \frac{92,3 \text{ g}}{12 \text{ g/mol}} = 7,69 \div 7,69 = 1$$

$$\text{Mol } H = \frac{7,7 \text{ g}}{1 \text{ g/mol}} = 7,7 \div 7,69 = 1$$



2 Na = 32,4%

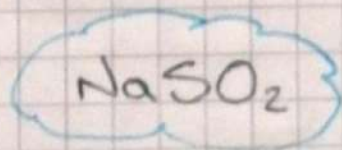
S = 22,5%

O = 45,1%

$$\text{Mol } Na = \frac{32,4 \text{ g}}{22 \text{ g/mol}} = 1,47 / 1,40 = 1$$

$$\text{Mol } S = \frac{22,5 \text{ g}}{16 \text{ g/mol}} = 1,40 / 1,40 = 1$$

$$\text{Mol } O = \frac{45,1 \text{ g}}{16 \text{ g/mol}} = 2,81 / 1,40 = 2$$



$$3 \quad C = 48\%$$

$$H = 4\%$$

$$N = 22,4\%$$

$$S = 12,8\% \text{ Azufre}$$

$$O = 12,8\%$$

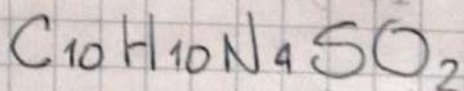
$$C = \frac{48 \text{ g}}{12 \text{ g/mol}} = 4 / 0,4 = 10$$

$$H = \frac{4 \text{ g}}{1 \text{ g/mol}} = 4 / 0,4 = 10$$

$$N = \frac{22,4 \text{ g}}{14 \text{ g/mol}} = 1,6 / 0,4 = 4$$

$$S = \frac{12,8 \text{ g}}{32 \text{ g/mol}} = 0,4 / 0,4 = 1$$

$$O = \frac{12,8 \text{ g}}{16 \text{ g/mol}} = 0,8 / 0,4 = 2$$

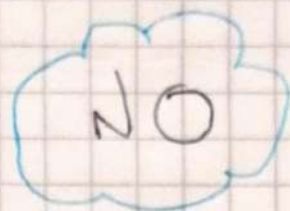


$$4 \quad N = 0,079 \text{ g}$$

$$O = 0,181 \text{ g}$$

$$N = \frac{0,079 \text{ g}}{14 \text{ g/mol}} = 0,0019,00$$

$$O = \frac{0,181 \text{ g}}{16 \text{ g/mol}} = 0,0113,00$$



$$5 \quad Na = 21,6\%$$

$$Cl = 33,3\%$$

$$O = 45,1\%$$

$$Na = \frac{21,6 \text{ g}}{22 \text{ g/mol}} = 0,9818,18 = 1$$

$$Cl = \frac{33,3 \text{ g}}{35 \text{ g/mol}} = 0,9514,29 = 1$$

$$O = \frac{45,1 \text{ g}}{16 \text{ g/mol}} = 2,8187,5 = 2$$

