



1 Calcula la masa molecular de las siguientes moléculas. ↓ ≠ ↑

a CaCO_3

$$\begin{aligned}\text{Ca} &= 40 \cdot 1 = 40 \\ \text{C} &= 12 \cdot 1 = 12 \\ \text{O} &= 16 \cdot 3 = 48 \\ &\hline &100\end{aligned}$$

b $\text{Fe(NO}_3)_3$

$$\begin{aligned}\text{Fe} &= 56 \cdot 1 = 56 \\ \text{N} &= 14 \cdot 3 = 42 \\ \text{O} &= 16 \cdot 9 = 144 \\ &\hline &242\end{aligned}$$

c HCl

$$\begin{aligned}\text{H} &= 1 \cdot 1 = 1 \\ \text{Cl} &= 35 \cdot 1 = 35 \\ &\hline &36\end{aligned}$$

d Al(OH)_3

$$\begin{aligned}\text{Al} &= 27 \cdot 1 = 27 \\ \text{O} &= 16 \cdot 3 = 48 \\ \text{H} &= 1 \cdot 3 = 3 \\ &\hline &78\end{aligned}$$

e HNO_3

$$\begin{aligned}\text{H} &= 1 \cdot 1 = 1 \\ \text{N} &= 14 \cdot 1 = 14 \\ \text{O} &= 16 \cdot 3 = 48 \\ &\hline &63\end{aligned}$$

f H_2SO_4

$$\begin{aligned}\text{H} &= 1 \cdot 2 = 2 \\ \text{S} &= 32 \cdot 1 = 32 \\ \text{O} &= 16 \cdot 4 = 64 \\ &\hline &98\end{aligned}$$

g $\text{C}_6\text{H}_{12}\text{O}_6$

$$\begin{aligned}\text{C} &= 12 \cdot 6 = 72 \\ \text{H} &= 1 \cdot 12 = 12 \\ \text{O} &= 16 \cdot 6 = 96 \\ &\hline &180\end{aligned}$$

h NaOH

$$\begin{aligned}\text{Na} &= 23 \cdot 1 = 23 \\ \text{O} &= 16 \cdot 1 = 16 \\ \text{H} &= 1 \cdot 1 = 1 \\ &\hline &40\end{aligned}$$

i MgO

$$\begin{aligned}\text{Mg} &= 24 \cdot 1 = 24 \\ \text{O} &= 16 \cdot 1 = 16 \\ &\hline &40\end{aligned}$$

j CuSO_4

$$\begin{aligned}\text{Cu} &= 63 \cdot 1 = 63 \\ \text{S} &= 32 \cdot 1 = 32 \\ \text{O} &= 16 \cdot 4 = 64 \\ &\hline &159\end{aligned}$$



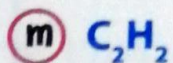
$$\text{N} = 14 \cdot 1 = 14$$

$$\text{H} = 1 \cdot 3 = \frac{3}{17}$$



$$\text{C} = 12 \cdot 6 = 72$$

$$\text{H} = 1 \cdot 14 = \frac{14}{86}$$



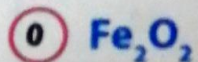
$$\text{C} = 12 \cdot 2 = 24$$

$$\text{H} = 1 \cdot 2 = \frac{2}{26}$$



$$\text{C} = 12 \cdot 1 = 12$$

$$\text{O} = 16 \cdot 2 = \frac{32}{44}$$



$$\text{Fe} = 56 \cdot 2 = 112$$

$$\text{O} = 16 \cdot 3 = \frac{48}{160}$$