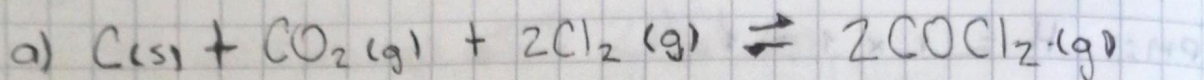
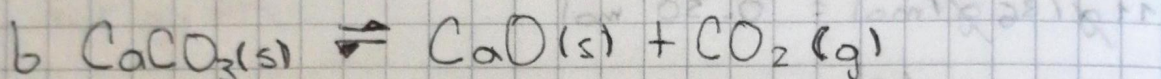


Ejemplos Constante de equilibrio para K_p :



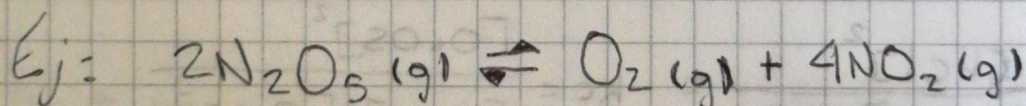
$$K_p = \frac{(\text{COCl}_2)^2}{\text{CO}_2 \cdot (\text{Cl}_2)^2}$$



$$K_p = \text{CO}_2$$

Unicamente se calculan los que son gaseosos
(g)

Presiones parciales



$$P_{\text{N}_2\text{O}_5} = 2.00 \text{ atm}$$

$$P_{\text{O}_2} = 0.296 \text{ atm}$$

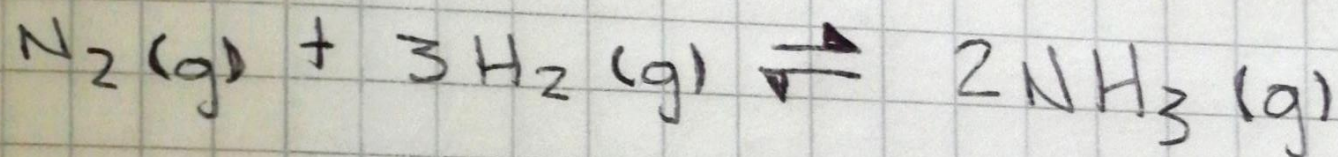
$$P_{\text{NO}_2} = 1.70 \text{ atm}$$

$$K_p = \frac{(\text{O}_2) (\text{NO}_2)^4}{(\text{N}_2\text{O}_5)^2}$$

Se reemplaza

$$K_p = \frac{(0.296)(1.70)^4}{(2.00)^2} = 0.618$$

Ejercicio



$$P_{\text{NH}_3} = 1.10 \text{ atm}$$

$$P_{\text{N}_2} = 0.37 \text{ atm}$$

$$P_{\text{H}_2} = 2.25 \text{ atm}$$

$$K_p = \frac{(1.10)^2}{(0.37)(2.25)} = 1.45345 \text{ atm}$$