

1 Calcula el pOH, las concentraciones si la solución tiene un pH de 9.6.

$$\text{pOH} = 14 - 9.6 = 4.4$$

$$\text{pH} = 9.6$$

$$[\text{H}^+] = 10^{-9.6} = \text{2.511} \times 10^{-10} \text{ M}$$

$$\text{OH}^- = 10^{4.4} = \cancel{25118.86} \quad 3.981 \times 10^{-5} \text{ M}$$

$$2 \quad p\text{OH} =$$

$$p\text{OH} = 2.50$$

$$p\text{OH} = -2.50 + 14 = p\text{H}$$

$$p\text{H} = 11.5$$

$$\text{H} = 3.162 \times 10^{-12} \text{ M}$$

$$\text{OH} = 0.003 \text{ M}$$

3

$$\text{H} = 2.4 \times 10^{-6} \text{ M}$$

$$p\text{H} = 2.4$$

$$p\text{OH} = -2.4 + 14 = 11.6$$

$$\text{OH} = 2.511 \times 10^{-12} \text{ M}$$

4

$$\text{OH} = 4.45 \times 10^{-6} \text{ M}$$

$$p\text{H} = -4.45 + 14 = 9.55 \text{ M}$$

$$\text{H} = 2.818 \times 10^{-10} \text{ M}$$

$$p\text{OH} = 4.45$$