

$$q = mc (t_f - t_i) \quad pQ = 55 (40^\circ\text{C} - 94^\circ\text{C})$$

$$Q = 55$$

$$Q = 55 \times 0.45 \text{ J/g}^\circ\text{C}$$

$$Q = 24.75 \times (40^\circ\text{C} - 94^\circ\text{C})$$

$$Q = 24.75 \times 54$$

$$Q = 1336.5 \text{ cal}$$

$$Q = mc (t_f - t_i) \quad pQ = 60 \text{ g } (100^\circ\text{C})$$

$$Q = 60 \text{ g } (100^\circ\text{C})$$

$$Q = 60 \text{ g} \times 540 \text{ cal/g}$$

$$Q = 60 \text{ g} \times 540 \text{ cal/g}$$

$$Q = 32.400 \text{ J}$$

$$Q = 32.400 \text{ J} \times 4.182 \text{ J}$$

$$Q = 135.400,98 \text{ J}$$

$$Q = 270 \text{ g} = (1 \text{ cal / g } ^\circ\text{C})$$

$$Q = 230 \text{ g} \times 1^\circ\text{C}$$

$$Q = 2,7^\circ\text{C} - 80^\circ\text{C}$$

$$Q = 77,3^\circ\text{C}$$

77,3°C se cedio para 730

calorias para calentar el otro material.