

Solucion

1

$$C \quad 92.3\% \div 12 = 7.691 \quad 7691 = 1$$

$$H \quad 7.7\% \div 1 = 7.777 \quad 7691 \div 7.777 = 1$$

→ CH
Resultado

2

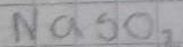
$$Na \quad 32.4\% \div 23 = 1.408 \quad 1406 \div 1406 = 1$$

$$S \quad 22.5\% \div 32 = 0.703 \quad 1406 \div 1406 = 1$$

$$O \quad 45.1\% \div 16 = 2.818 \quad 2818 \div 1406 = 2$$

Resultado

↓



3

$$C \quad 48\% \div 12$$

$$H \quad 4\% \div 1$$

$$N \quad 22.4\% \div 14$$

$$\text{S} \quad \text{12.8\%} \div 32$$

$$\text{O} \quad \text{12.8\%} \div 16$$

4

$$\text{N} \quad 0.079 \div 14 = 0.005 + 0.005 = 3.6$$

$$\text{O} \quad 0.181 \div 16 = 0.011 \div 0.011 = 1$$

N_3O_1

5

$$\text{Na} \quad 21.6\% \div 23 = 0.939 \div 0.939 = 1$$

$$\text{C} \quad 33.3\% \div 12 = 2.775 \div 0.939 = 2$$

$$\text{O} \quad 45.1\% \div 16 = 2.818 \div 0.939 = 3$$

$\text{Na}_1\text{C}_2\text{O}_3$

solucion

1). $H = 3 \times 1 = 3 \div 98 = 0.030 \times 100 = 3\%$

$P = 1 \times 31 = 31 \div 98 = 0.316 \times 100 = 31.6\%$

$O = 4 \times 16 = \frac{64}{98 \text{ g mol}} \div 98 = 0.653 \times 100 = 65.3\%$
99.9%

2).

$Pb = 1 \times 207 = 207 / 275 = 0.752 \times 100 = 75.2\%$

$O = 4 \times 16 = 64 / 275 = 0.232 \times 100 = 23.2\%$

$H = 4 \times 1 = \frac{4}{275 \text{ g mol}} \div 275 = 0.014 \times 100 = 1.4\%$
99.8%

3).

$Ni = 2 \times 58 = 116 / 296 = 0.391 \times 100 = 39.1\%$

$O = 3 \times 12 = 36 / 296 = 0.121 \times 100 = 12.1\%$

$O = 9 \times 16 = \frac{144}{296 \text{ g mol}} \div 296 = 0.486 \times 100 = 48.6\%$
99.8%

4)

$$H = 2 \times 1 = 2 \quad 98 = 0.020 \times 100 = 2.0\%$$

$$S = 1 \times 32 = 32 \quad 98 = 0.326 \times 100 = 32.6\%$$

$$O = 4 \times 16 = 64 \quad 98 = 0.653 \times 100 = \frac{65.3\%}{99.9}$$

$\frac{98 \text{ g}}{\text{mol}}$

5)

$$H = 2 \times 1 = 2 \quad 18 = 0.111 \times 100 = 11.1\%$$

$$O = 1 \times 16 = 16 \quad 18 = 0.888 \times 100 = \frac{88.8\%}{99.9}$$

$\frac{18 \text{ g}}{\text{mol}}$