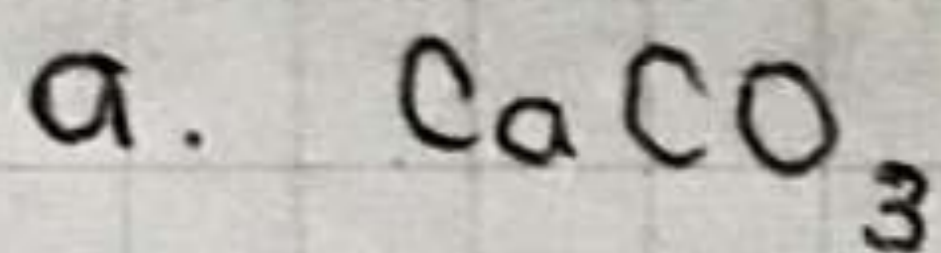


ESTEQUIOMETRIA

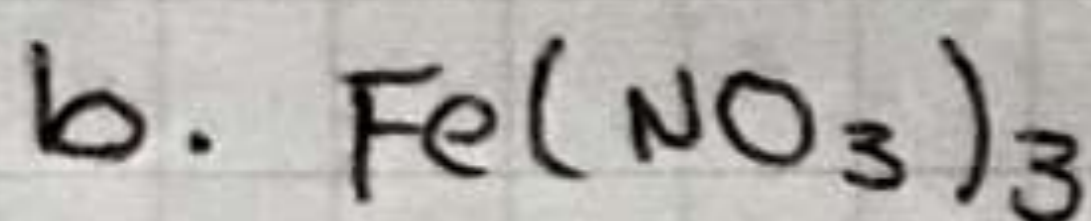
Analizar y comprender los procedimientos estequiométricos.

u.m.a = Unidad
Masa
Atómica

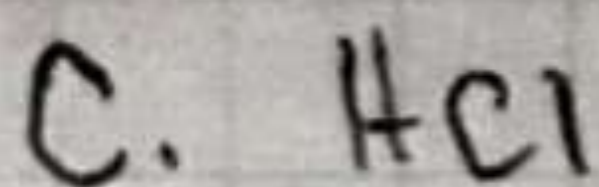
EJERCICIOS



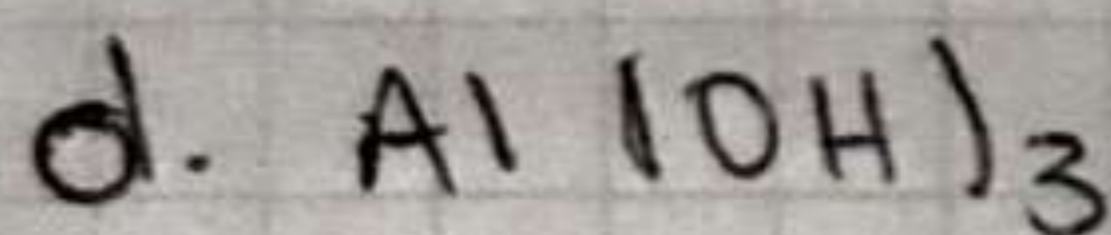
$$\begin{aligned}\text{Ca} &= 40 \cdot 1 = 40 \text{ u.m.a.} \\ \text{C} &= 12 \cdot 1 = 12 \text{ u.m.a.} \\ \text{O} &= 16 \cdot 3 = \underline{48 \text{ u.m.a.}} \\ &100 \text{ u.m.a.}\end{aligned}$$



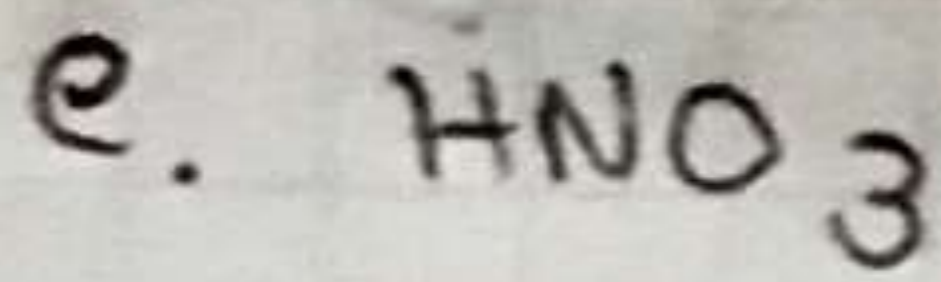
$$\begin{aligned}\text{Fe} &= 56 \cdot 1 = 56 \text{ u.m.a.} \\ \text{N} &= 14 \cdot 3 = 42 \text{ u.m.a.} \\ \text{O} &= 16 \cdot 9 = 144 \text{ u.m.a.}\end{aligned}$$



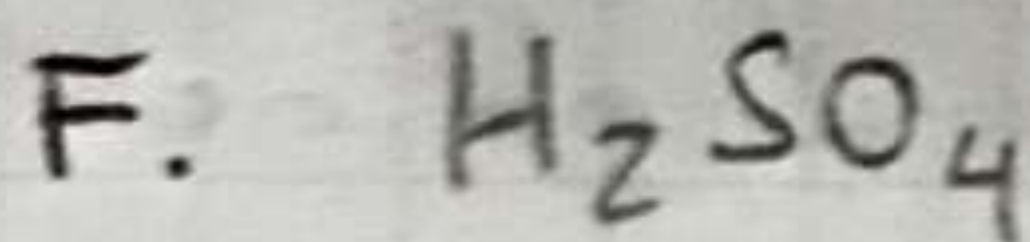
$$\begin{aligned}\text{H} &= 1 \cdot 1 = 1 \text{ u.m.a.} \\ \text{Cl} &= 35 \cdot 1 = \underline{35 \text{ u.m.a.}} \\ &36 \text{ u.m.a.}\end{aligned}$$



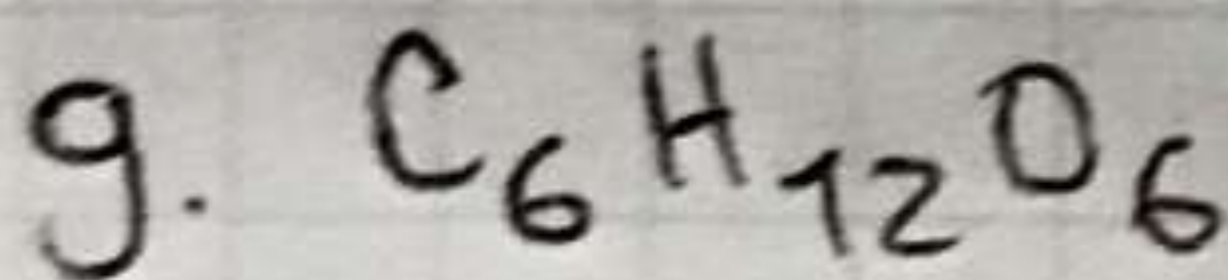
$$\begin{aligned}\text{Al} &: 27 \cdot 1 = 27 \text{ u.m.a.} \\ \text{H} &: 1 \cdot 3 = 3 \text{ u.m.a.} \\ \text{O} &: 16 \cdot 3 = \underline{48 \text{ u.m.a.}} \\ &78 \text{ u.m.a.}\end{aligned}$$



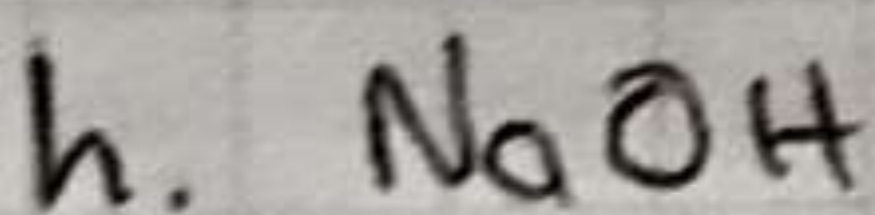
$$\begin{aligned} \text{H} &= 1 \cdot 1 = 1 \text{ uma} \\ \text{N} &= 14 \cdot 1 = 14 \text{ uma} \\ \text{O} &= 16 \cdot 3 = \underline{48 \text{ uma}} \\ &50 \text{ uma} \end{aligned}$$



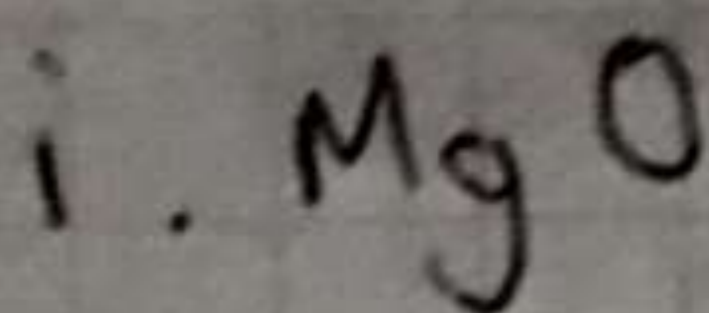
$$\begin{aligned} \text{H} &= 1 \cdot 2 = 2 \text{ uma} \\ \text{S} &= 32 \cdot 1 = 32 \text{ uma} \\ \text{O} &= 16 \cdot 4 = \underline{64 \text{ uma}} \\ &98 \text{ uma} \end{aligned}$$



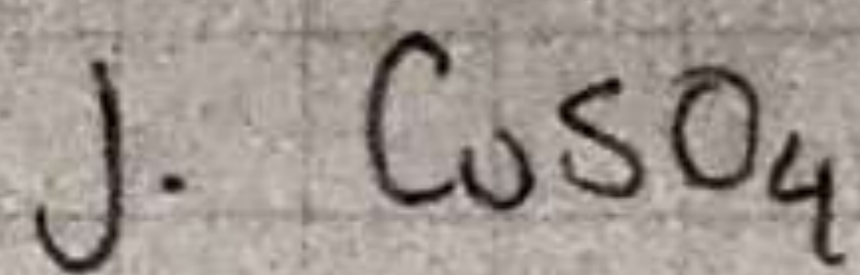
$$\begin{aligned} \text{C} &= 12 \cdot 6 = 72 \text{ uma} \\ \text{H} &= 1 \cdot 12 = 12 \text{ uma} \\ \text{O} &= 16 \cdot 6 = \underline{96 \text{ uma}} \\ &180 \text{ uma} \end{aligned}$$



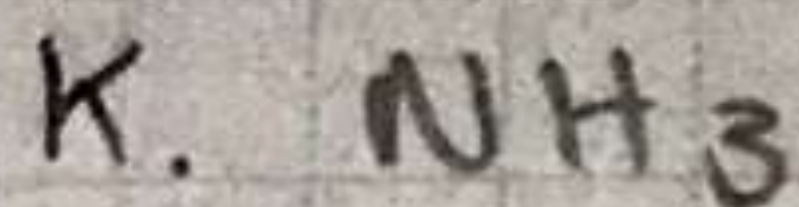
$$\begin{aligned} \text{Na} &= 23 \cdot 1 = 23 \text{ uma} \\ \text{O} &= 16 \cdot 1 = 16 \text{ uma} \\ \text{H} &= 1 \cdot 1 = \underline{1 \text{ uma}} \\ &40 \text{ uma} \end{aligned}$$



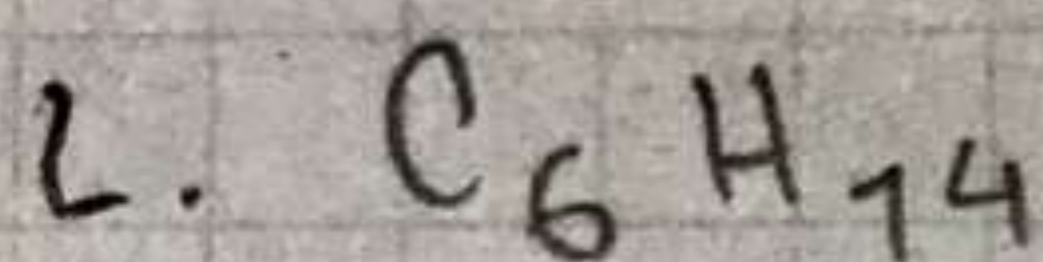
$$\begin{aligned} \text{Mg} &= 24 \cdot 1 = 24 \text{ uma} \\ \text{O} &= 16 \cdot 1 = \underline{16 \text{ uma}} \\ &40 \text{ uma} \end{aligned}$$



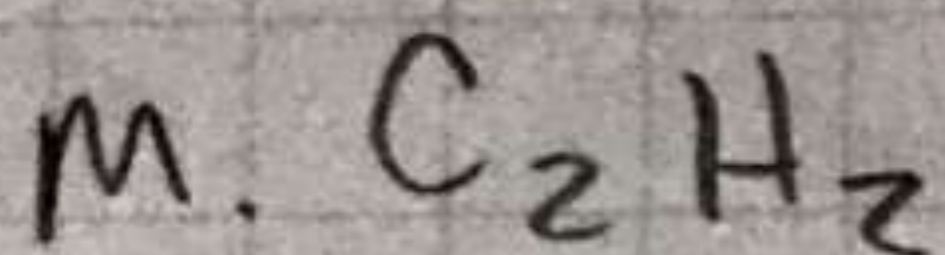
$$\begin{aligned}\text{Cu} &: 63 \cdot 1 = 63 \text{ uma} \\ \text{S} &: 32 \cdot 1 = 32 \text{ uma} \\ \text{O} &: 16 \cdot 4 = \underline{64 \text{ uma}} \\ & 159 \text{ uma}\end{aligned}$$



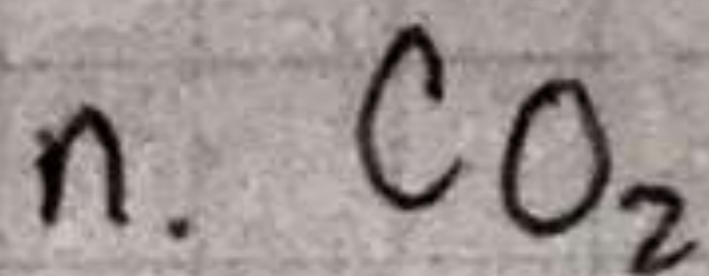
$$\begin{aligned}\text{N} &: 14 \cdot 1 = 14 \text{ uma} \\ \text{H} &: 1 \cdot 3 = \underline{3 \text{ uma}} \\ & 17 \text{ uma}\end{aligned}$$



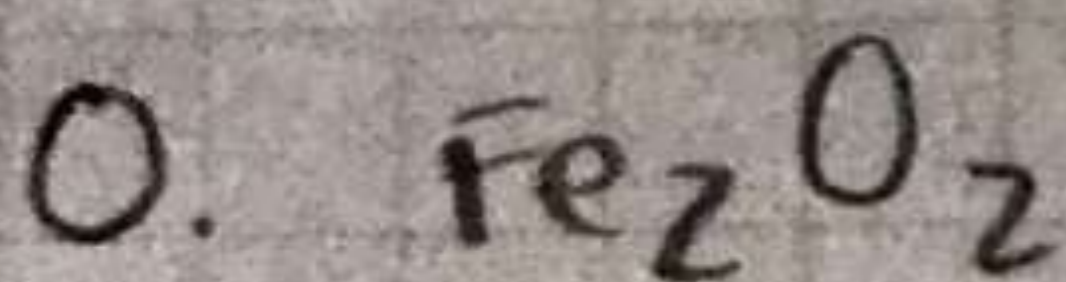
$$\begin{aligned}\text{C} &: 12 \cdot 6 = 72 \text{ uma} \\ \text{H} &: 1 \cdot 14 = \underline{14 \text{ uma}} \\ & 86 \text{ uma.}\end{aligned}$$



$$\begin{aligned}\text{C} &: 12 \cdot 2 = 24 \text{ uma} \\ \text{H} &: 1 \cdot 2 = \underline{2 \text{ uma}} \\ & 26 \text{ uma.}\end{aligned}$$



$$\begin{aligned}\text{C} &: 12 \cdot 1 = 12 \text{ uma} \\ \text{O} &: 16 \cdot 2 = \underline{32 \text{ uma}} \\ & 44 \text{ uma}\end{aligned}$$



$$\begin{aligned}\text{Fe} &: 56 \cdot 2 = 112 \text{ uma} \\ \text{O} &: 16 \cdot 3 = \underline{48 \text{ uma}} \\ & 160 \text{ uma.}\end{aligned}$$