

~~CH = 5.351~~
 Ph = ~~5.351~~ - 5.351 + 14
 Ph = ~~8.692~~ 8.692 m
 H = ~~2.290~~ 2.290 x 10⁻⁹
 Pch 5.351

revisar

Identificar y comprender el procedimiento para obtener la constante de equilibrio

Constante de Equilibrio

$$K = \frac{C^c \cdot D^d}{A^a \cdot B^b}$$

Productos
Reactivos

K = Constante de equilibrio

A B = Productos

C D = Reactivos

[A] = Concentración de A en moles
 a = Número de moles de A

~~log p = 5~~ Salt = 5
 log p = 5
 p = 10⁵

N₂ O₄ p

Homos = lo mismo no se diferencia
 hetero = si se genera diferencia

Es la diferencia heterogénea

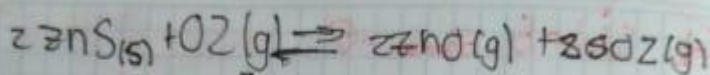


$$K = \frac{[ZnH_3] + 4[H_2O]^2}{[ZnO_2] + 7[H_2]}$$

$$K_{eq} = \frac{[p]^2}{[r]}$$

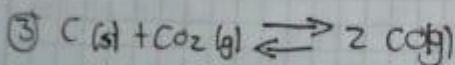
$$K_c = \frac{[NH_3]^2}{[NO_2]^2} \times \frac{[H_2O]^4}{[H_2]^7}$$

Heterogénea



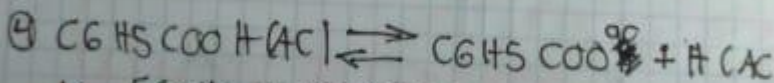
$$K_c = \frac{[\text{ZnO}]^2 \times [\text{SO}_2]^2}{[\text{ZnS}]^2 \times [\text{O}_2]}$$

Heterogonia



$$K_c = \frac{[\text{CO}]^2}{[\text{CO}_2]}$$

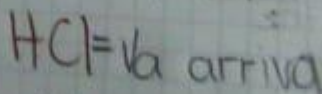
Heterogonia



$$K_c = \frac{[\text{C}_6\text{H}_5\text{COO}^-] \times [\text{H}^+]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

Homogonia

$$g/pm \Rightarrow m/L$$



37g

71g

36.5g

3.08

0.058

0.0508

$$\frac{(0.0508)^2}{3.08 \cdot 0.058}$$

$$K_c = 0.13$$