



LUN



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## Composición porcentual

### Ejercicios

Calcular la composición porcentual de:

1  $H_3PO_4$ : Ácido fosfórico

$$H = 1 \times 3 = 3 \text{ ———— } / 98 = 0.03 \times 100 = 3\%$$

$$P = 31 \times 1 = 31 \text{ ———— } / 98 = 0.31 \times 100 = 31\%$$

$$O = 16 \times 4 = \frac{64}{98 \text{ g/mol}} \text{ ———— } / 98 = 0.65 \times 100 = \frac{65\%}{99}$$

2  $Pb(OH)_4$ : Hidróxido de plomo (IV)

$$Pb = 207 \times 1 = 207 \text{ ———— } / 272 = 0.76 \times 100 = 76\%$$

$$O = 16 \times 4 = 64 \text{ ———— } / 272 = 0.23 \times 100 = 23\%$$

$$H = 1 \times 4 = \frac{4}{272 \text{ g/mol}} \text{ ———— } / 272 = 0.01 \times 100 = \frac{1\%}{100}$$

3  $Ni_2(CO_3)_3$ : Carbonato de níquel (II)

$$Ni = 58 \times 2 = 116 \text{ ———— } / 161 = 0.72 \times 100 = 72\%$$

$$C = 12 \times 3 = 36 \text{ ———— } / 161 = 0.22 \times 100 = 22\%$$

$$O = 16 \times 9 = \frac{144}{161 \text{ g/mol}} \text{ ———— } / 161 = 0.05 \times 100 = \frac{5\%}{99}$$



4  $H_2SO_4$ : Acido sulfurico

$$H = 1 \times 2 = 2 \quad \text{---} \quad 98 = 0.02 \times 100 = 2\%$$

$$S = 32 \times 1 = 32 \quad \text{---} \quad 98 = 0.32 \times 100 = 32\%$$

$$O = 16 \times 4 = \frac{64}{98 \text{ g/mol}} \quad \text{---} \quad 98 = 0.65 \times 100 = \frac{65\%}{99}$$

5  $H_2O$ : Agua

$$H = 1 \times 2 = 2 \quad \text{---} \quad 18 = 0.11 \times 100 = 11\%$$

$$O = 16 \times 1 = \frac{16}{18 \text{ g/mol}} \quad \text{---} \quad 18 = 0.88 \times 100 = \frac{88\%}{99}$$