


$$V_{\text{sólido}} = 55 \text{ gr}$$

$$V_{\text{solvente}} = 239 \text{ gr}$$

$$V_{\text{solución}} = 294 \text{ gr}$$

$$\frac{55}{294} = 0,187 \times 100 = 18,70\%$$

g



$$V_{\text{sólido}} = 2,3 \text{ K}$$

$$V_{\text{solvente}} = 6 \text{ L}$$

$$V_{\text{solución}} = 8,3$$

$$\frac{2,3}{8,3} = 0,277 \times 100 = 27,7\%$$

~~80~~

## Ejercicios

$$\textcircled{1} \quad M_{\text{solute}} = 13\text{g}$$

$$M_{\text{solvente}} = \underline{97\text{g}}$$

$$M_{\text{solución}} = 110\text{g}$$

$$13 \div 110 = 0,118 \times 100 = 11,818 \%$$

$$\textcircled{2} \quad M_{\text{solute}} = 10\text{g}$$

$$M_{\text{solvente}} = \underline{113\text{g}}$$

$$M_{\text{solución}} = 123\text{g}$$

$$10 \div 123 = 0,081 \times 100 = 8,130 \%$$

$$\textcircled{3}$$

$$V_{\text{solute}} = 2,04 \text{ ml}$$

$$V_{\text{solvente}} = \underline{15 \text{ ml}}$$

$$V_{\text{solución}} = 17,04$$

$$2,04 \div 17,04 = 0,119 \times 100 = 11,977 \%$$

$$\textcircled{a} \quad V_{\text{Soluta}} = 2,09 \text{ ml}$$

$$V_{\text{Solvente}} = \underline{11 \text{ ml}}$$

$$V_{\text{Soluci3n}} = 13,09$$

$$2,09 \div 13,09 = 0,159 \quad \times 100 = 15,966 \%$$