

1.2



= iónico



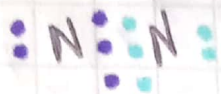
F = 3.99

= 2.38

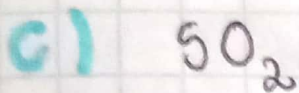
Be = 1.57



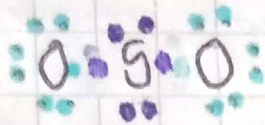
= covalente no polar



N = 3.04



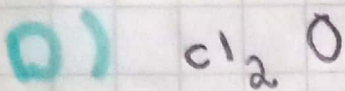
= covalente polar



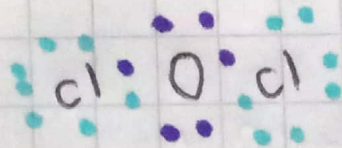
O = 3.44

= 0.86

S = 2.58



= iónico

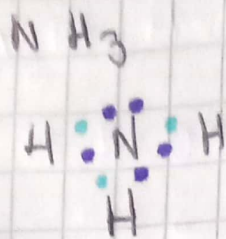


Cl = 3.16

= 0.28

O = 3.44

e)



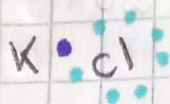
= covalente polar

\bullet N = 3.04 \bullet 0.84
 \bullet H = 2.2

f)

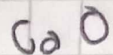


= ionico

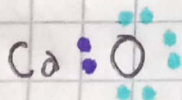


\bullet K = 0.82 \bullet 2.34
 \bullet cl = 3.16

g)



= ionico

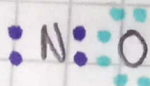


\bullet Ca = 1 \bullet 2.44
 \bullet O = 3.44

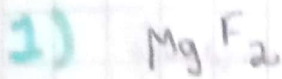
h)



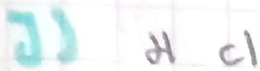
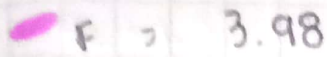
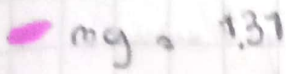
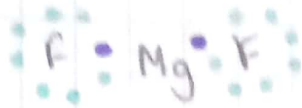
= covalente no polar



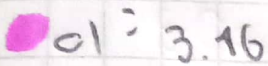
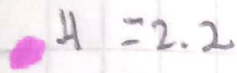
\bullet N = 3.04 \bullet 0.4
 \bullet O = 3.44



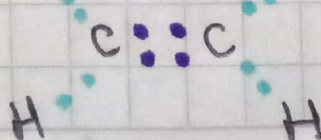
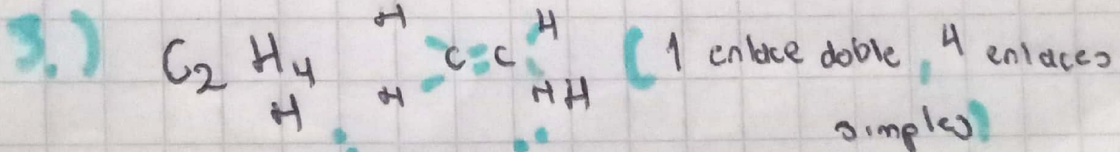
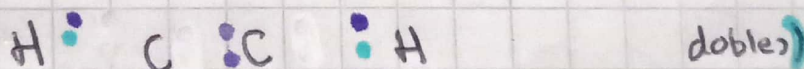
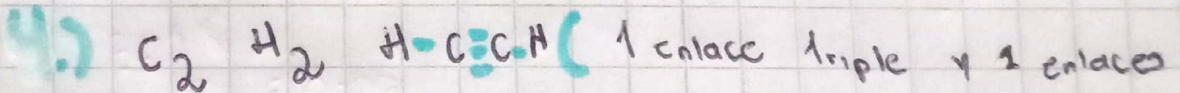
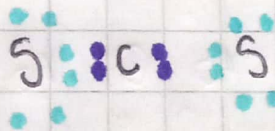
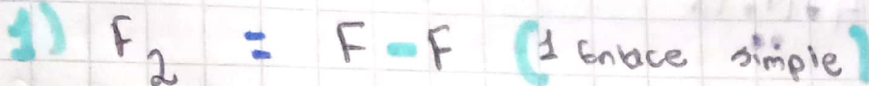
= Iónico

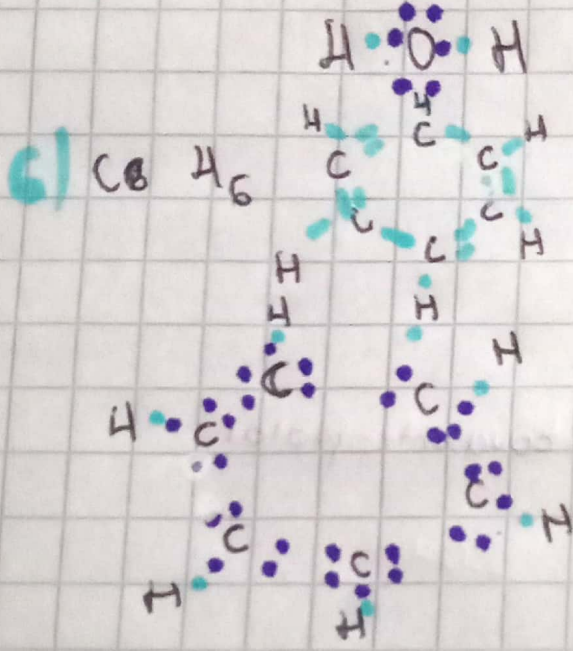
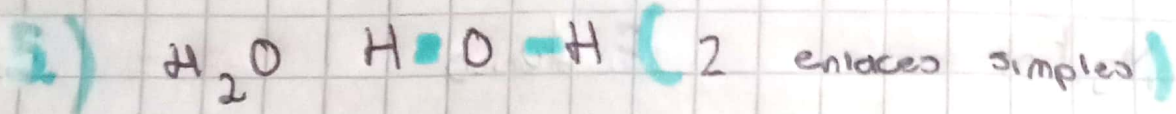


= covalente polar

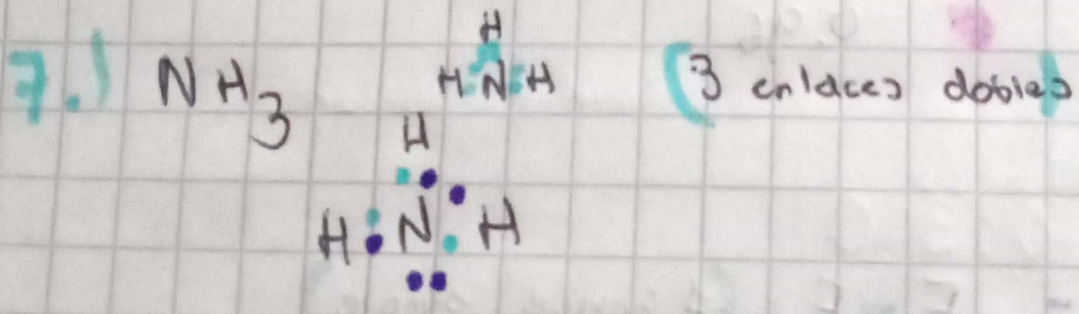


2.)





(8 enlaces simples y 3 enlaces dobles)



(3 enlaces dobles)