

Decanolo

I =

$$C = 92,3\%$$

$$H = 7,7\%$$

Mol C

Mol H

$$\frac{7,7 \text{ g}}{1 \text{ g/mol}}$$

$$\frac{92,3 \text{ g}}{12 \text{ g/mol}}$$

$$= 7,691$$

$$= 7,7$$

$$\text{Mol} = 7,691 / 7,7 = 0$$

$$\text{Mol} = 7,7 / 7,7 = 1$$

C H

$$2 = \text{Na} = 32,4\% \quad \text{O} = 45,7\%$$

$$\text{S} = 22,5\%$$

Mol

Na

$$\frac{32,4 \text{ g}}{23 \text{ g/mol}}$$

$$1,408$$

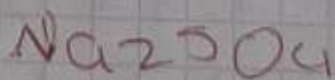
$$\text{Mol } 5 \quad \frac{22.59}{32 \text{ g/mol}} \quad 0.703$$

$$\text{Mol } 0 \quad \frac{43.79}{16 \text{ g/mol}} \quad 2.848$$

$$\text{mol } 1.408 / 0.703 = 2$$

$$\text{mol } 0.703 / 0.703 = 1$$

$$\text{mol } 2.848 / 0.703 = 4$$

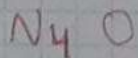


$$4 \quad N = 0.079\% \\ O = 0.181\%$$

$$\text{Mol } N \quad \frac{0.079 \text{ g}}{14 \text{ g/mol}} = 5.642$$

$$\text{Mol } O \quad \frac{0.181 \text{ g}}{16 \text{ g/mol}} = 0.011$$

$$\text{Mol } \quad 5.642 / 0.011 \quad 4 \\ 0.011 / 0.011 = 1$$



$$5 \quad Na = 21.6\% \quad O = 45.1\% \\ Cl = 33.3\%$$

$$\text{Mol } Na \quad \frac{21.6 \text{ g}}{23 \text{ g/mol}} = 0.939$$

$$\text{Mol } Cl \quad \frac{33.3 \text{ g}}{35 \text{ g/mol}} = 0.951$$

$$\begin{array}{r} \text{Mol } 0 \\ \hline 43.7 \text{ g} \\ 16 \text{ g/mol} \\ \hline 2.818 \end{array}$$

$$\text{Mol } 0.939 / 0.939 \quad 1$$

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$$\text{Mol } 2.818 / 0.939 \quad 3$$

$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$