

Ejercicios de molaridad.

Solución 1 =

$$\text{HNO}_3 \quad m = \frac{\text{gramos}}{(\text{Pm})(\text{kg})}$$

$$\text{masa} = 95 \text{ g}$$

$$\text{volumen} = 0,025$$

$$\text{Pm} = 63$$

$$\text{H} = 1 \times 1 = 1$$

$$\text{N} = 1 \times 14 = 14$$

$$\text{O} = 3 \times 16 = \underline{48}$$
$$63$$

$$m = \frac{95}{63 \cdot 0,025} = \frac{95}{1,575} =$$
$$= \textcircled{60,317 \text{ m}}$$

2



$$\text{masa} = 95 \text{ g}$$

$$\text{volumen} = 0,025$$

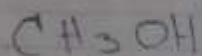
$$m = \frac{\text{m}}{\text{kg}}$$

$$\begin{aligned}
 H &= 7 \times 1 = 7 \\
 N &= 1 \times 14 = 14 \\
 O &= 3 \times 16 = \frac{48}{63}
 \end{aligned}$$

$$\frac{92g}{0.2g \text{ mol}} = 1.207 \text{ moles}$$

$$\frac{1.207 \text{ mol}}{0.027} = \text{60.317 mol}$$

Segundo punto =



$$\text{masa} = 12g$$

m gramos
(P m)(Kg)

$$\text{Volumen} = 0,05$$

$$P_m = 32$$

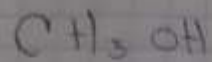
$$C = 1 \times 12 = 12$$

$$H = 3 \times 1 = 3$$

$$H = 1 \times 1 = \frac{1}{32}$$

$$O = 1 \times 16 = 16$$

$$m = \frac{15}{(32)(0.05)} = \frac{15}{16} = \text{9,375 mol}$$



$$mass = 10g$$

$$Volume = 0,03$$

$$P_m = 32$$

$$m = \frac{m}{kg}$$

$$C = 1 \times 12 = 12$$

$$H = 3 \times 1 = 3$$

$$O = 1 \times 16 = 16$$

$$H = 1 \times 1 = \frac{1}{32}$$

$$\frac{10g}{32g/mol} = 0,468 \text{ moles}$$

$$\frac{0,468}{0,03} = 9,375 \text{ m}$$