DD MM AA

Desarrollo !

- C = 40 / 12 g mol = 3 mol / 3 = 1
  H = 6.7 / 1 g mol = 6.7 mol / 3 = 2
  O = 53.3 / 16 gmol = 3.331 / 3 = 1

Formula empirica: CH_2O

C = 12 x 1 = 12 H = 1 x 2 = 1 O = 16 x 1 = 16 — 90 / 29 = 3

- **Formula molecular:** $C_3H_6O_3$

- C = 37,8 / 12 g/mol = 3.15 mol / 1,5 = 2
  H = 6.3 / 1 g/mol = 6.3 mol / 1,5 = 4
  Cl = 55.8 / 35 g/mol = 1.57 mol / 1,5 = 1

Formula empirica: C_2H_2Cl

C = 12 x 2 = 24
H = 1 x 4 = 4
Cl = 35 x 1 = 35
63

127 g / 63 = 2

2 x 2 = 4
2 x 4 = 8
2 x 1 = 2

Formula molecular: $C_4H_8Cl_2$