

Solución.

- $H_2 Pb_6$

$$H = 3 \times 1 = 3 \text{ g} = 0.030 \times 100 = 3\%$$

$$Pb = 1 \times 31 = 31 \text{ g} = 0.316 \times 100 = 31.6\%$$

$$O = 4 \times 16 = 64 \text{ g} = 0.653 \times 100 = 65.3\%$$

98 g/mol

- $Pb(OH)_4$

$$Pb = 1 \times 207 = 207 \div 275 = 0.752 \times 100 = 75.2\%$$

$$O = 4 \times 16 = 64 \div 275 = 0.232 \times 100 = 23.2\%$$

$$H = 4 \times 1 = 4 \div 275 = 0.014 \times 100 = 1.4\%$$

98.8

- $Ni_2(CO_3)_3$

$$Ni = 2 \times 58 = 116 \div 296 = 0.391 \times 100 = 39.1\%$$

$$C = 3 \times 12 = 36 \div 296 = 0.121 \times 100 = 12.1\%$$

$$O = 9 \times 16 = 144 \div 296 = 0.486 \times 100 = 48.6\%$$

99.8

- $H_2 SO_4$

$$H = 2 \times 1 = 2 \div 98 = 0.020 \times 100 = 2\%$$

$$S = 1 \times 32 = 32 \div 98 = 0.326 \times 100 = 32.6\%$$

$$O = 4 \times 16 = 64 \div 98 = 0.653 \times 100 = 65.3\%$$

99.9



DD

MM

AA

• H₂O

$$H = 2 \times 1 = 2 \div 3 = 0.666 \times 100 = 66.6\%$$

$$O = 1 \times 1 = \frac{1}{3 \text{ g/mol}} \div 3 = 0.333 \times 100 = \frac{33.3\%}{99.9}$$