

Unidades químicas de concentración

molaridad

$$M = \frac{\text{moles de soluto}}{\text{litros de solución}}$$

molaridad

$$m = \frac{\text{moles de soluto}}{\text{kg de solvente}}$$

normalidad

$$N = \frac{\text{equivalente gr soluto}}{\text{litro de solución}}$$

Calcular

$$12 \times 2 \times 10 = 24$$

$$46 \times 1 = 46 \quad +$$

$$0.1 \times 16 = \frac{16}{46g}$$

$$82.5 + 96 = 1.795 \text{ moles} = \frac{1.795 \text{ mol}}{0.4506}$$

$$= 3.984 M$$

$$= 4 M$$

scribble

Example 3:

$$m = 92.5$$

$$P_m = 46$$

$$V = 0.450$$

$$\frac{92.5}{(46)(0.450)}$$

\sqrt{x}

$$\downarrow$$

$$20.7$$

$$\frac{92.5}{20.7} = 3.985 \text{ M} \rightarrow \text{molaridad}$$

$$= 4$$

2 Punto

$$n = 4.70$$

$$V = 7000 \text{ ml} \div 1000 = 7 \text{ L}$$

$$\text{Por 1} = \frac{4.70 \text{ mol}}{7 \text{ L}} = 0.682 \text{ M molaridad}$$

3 Punto

$$\text{Por 1} = n = 3.65$$

$$3.65 \div 36 = 0.101$$

$$S = 2$$

$$\frac{0.101}{2} = 0.0505 \text{ M}$$

$$P_m \text{ H } 1 \times 1 = 1$$

\downarrow
molaridad

$$\text{Cl } 1 \times 35 = 35 \rightarrow 36$$

Formula 2

$$M = 3,66$$

$$B_1 = 36$$

$$\sqrt{=} 2$$

$$\begin{array}{r} 3,66 \\ \times 10 \\ \hline 70 \end{array}$$

$$\frac{36,9}{10} = 0,0369$$

4 Punto

For 1

$$m = 48,04$$

$$1 \times 1 \times 1 = 2$$

$$25 \times 32 = 32$$

$$0,4 \times 16 = 64$$

$$\frac{64}{88} \rightarrow \text{For}$$

$$n = 48,04 \div 88 = 0,50 \text{ mol}$$

$$\frac{0,500}{0,25} = 2 \text{ M}$$

↓
molaridad

For 2

$$M = 48,04$$

$$B_1 = 96$$

$$\sqrt{=} 0,25$$

$$\begin{array}{r} 48,04 \\ \times 0,25 \\ \hline 24,5 \end{array}$$

$$\frac{48,04}{24,5} = 2 \text{ M}$$

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molaridad