Formula Empírica C = 92.3% H = 7.7%

Mol C $\frac{92.3\ g}{12\ g/mol}$ = 7.69 mol / 7.69 = 1

= CH

Mol H $\frac{7.7\ g}{1\ g/mol}$ = 7.7 mol / 7.69 = 1

Na = 32.4% S = 22.5% O = 45.1%

Mol Na $\frac{32.4\ g}{22.98\ g/mol}$ = 1.40 mol / 0.7 = 2

Mol S $\frac{22.5\ g}{32.06\ g/mol}$ = 0.7 mol / 0.7 = 1 = Na_2SO_4

Mol O $\frac{45.1\ g}{16\ g/mol}$ = 2.8 mol / 0.7 = 4

C = 48% H = 4% N = 22.4% S = 12.8% O = 12.8%

Mol C $\frac{48}{12}$ = 4 mol / 0,4 = 10

Mol H $\frac{4}{1}$ = 4 mol / 0,4 = 10 = $C_{10}H_{10}N_4SO_2$

Mol N $\frac{22}{14}$ = 1.6 mol / 0.4 = 4

Mol S $\frac{12.8}{32}$ = 0,4 mol / 0,4 = 1

Mol O $\frac{12.8}{16}$ = 0,8 mol / 0,4 = 2

N = 0,079 g O = 0,181 g = NO_2

Mol N = $\frac{0,079}{14}$ = 0,005 mol / 0,005 = 1

Mol O = $\frac{0,181}{16}$ = 0,011 mol / 0,005 = 2.2

Na = 21.6% Cl = 33.3% O = 45.1%

Mol Na = $\frac{21.6}{23}$ = 0.93 mol / 0.93 = 1

Mol Cl = $\frac{33.3}{35}$ = 0.95 mol / 0.93 = 1

mol O = $\frac{45.1}{16}$ = 2.81 mol / 0.93 = 3

= $NaClO_3$