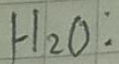
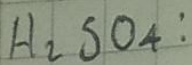
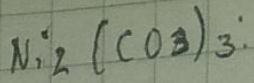
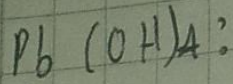
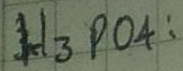


Calcular

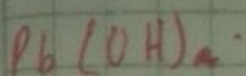


Solución

$$H: 1 \cdot 3 = 3 \div 98 = 0.030 \cdot 100 = 3.0\% H$$

$$P: 31 \cdot 1 = 31 \div 98 = 0.316 \cdot 100 = 31.6\% P$$

$$O: 16 \cdot 4 = \frac{64}{98 \text{ g/mol}} \div 98 = 0.653 \cdot 100 = 65.3\% O$$

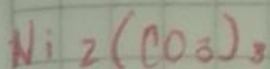


$$Pb: 207 \cdot 1 = 207 \div 275 = 0.752 \cdot 100 = 75.2 Pb$$

$$O: 16 \cdot 4 = 64 \div 275 = 0.232 \cdot 100 = 23.2 O$$

$$H: 1 \cdot 4 = 4 \div 275 = 0.014 \cdot 100 = 1.4 H$$

$275 g/mol$

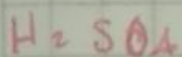


$$Ni: 59 \cdot 2 = 118 \div 222 = 0.531 \cdot 100 = 53.1 Ni$$

$$C: 12 \cdot 3 = 36 \div 222 = 0.162 \cdot 100 = 16.2 C$$

$$O: 16 \cdot 9 = 144 \div 222 = 0.648 \cdot 100 = 64.8 O$$

$222 g/mol$



$$H: 1 \cdot 2 = 2 \div 98 = 0.020 \cdot 100 = 2.0 H$$

$$S: 32 \cdot 1 = 32 \div 98 = 0.326 \cdot 100 = 32.6 S$$

$$O: 16 \cdot 4 = 64 \div 98 = 0.653 \cdot 100 = 65.3 O$$

$98 g/mol$



$$H: 1 \cdot 2 = 2 \div 18 = 0.111 \cdot 100 = 11.1 H$$

$$O: 16 \cdot 1 = 16 \div 18 = 0.888 \cdot 100 = 88.8 O$$

$18 g/mol$