

propósito identificar comprender el procedimiento para calcular la constante de equilibrio

$$K_c = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

K = constante

A, B = productos

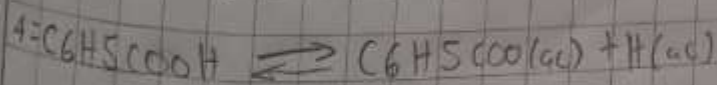
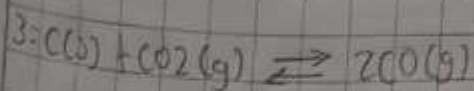
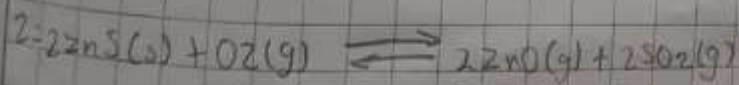
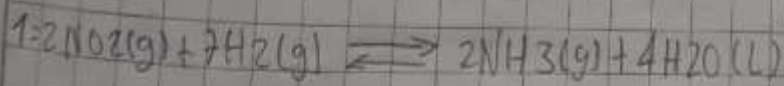
C, D = reactivos

[A] concentración de equilibrio de A en moles

a = número de moles de A

puntos importantes

cuando se puede llevar una constante de equilibrio



$$1: K_c = \frac{(\text{NH}_3)^2 \times (\text{H}_2\text{O})^4}{(\text{NO}_2)^2 \times (\text{H}_2)^7}$$

Heterogénea

$$K_c = \frac{(ZnO)_2 \times (SO_2)_2}{(ZnS)_2 \times (CO_2)_2}$$

Heterogeneous

$$K_c = \frac{(CO)_2}{(C) \times (CO_2)}$$

Heterogeneous

$$K_c = \frac{(C_6H_5COO^-) \times (H^+)}{(C_6H_5COOH)}$$

homogeneous

$$K_c = \frac{[NH_3]^2}{[NO_2][H_2]}$$

$$K_p = \frac{P_{NH_3}^2}{P_{NO_2} P_{H_2}}$$

$$K_c = \frac{[ZnO]^2 [SO_2]^2}{[O_2]}$$

$$K_p = \frac{P_{ZnO}^2 P_{SO_2}^2}{P_{O_2}}$$

$$K_c = \frac{[CO]^2}{[O_2]}$$

$$K_p = \frac{P_{CO}^2}{P_{O_2}}$$

$$n = g / PM$$

$$SO_2 = 120g$$

$$PM: S = 32 \times 1 = 32$$

$$O = 16 \times 2 = 32$$

$$n = g / PM$$

$$64g/mol$$

$$120g / 64g/mol = 1.875 mol$$

$$M = \text{mole} / L$$

$$M = 1.875 / 2L$$

$$M = 0.95$$