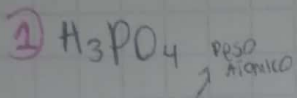


## Solución

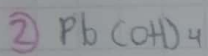


$$H = 1 \times 3 = 3 \div 98 = 0,030 \times 100 = 3\%$$

$$P = 31 \times 1 = 31 \div 98 = 0,316 \times 100 = 31,6\%$$

$$O = 16 \times 4 = \frac{64}{98} \div 98 = 0,653 \times 100 = 65,3\%$$

98 g/mol

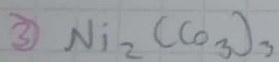


$$Pb = 1 \times 207 = 207 \div 275 = 0,752 \times 100 = 75,2\%$$

$$O = 4 \times 16 = 64 \div 275 = 0,232 \times 100 = 23,2\%$$

$$H = 4 \times 1 = \frac{4}{275} \div 275 = 0,014 \times 100 = 1,4\%$$

275 g/mol

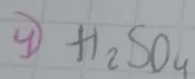


$$Ni = 2 \times 59 = 118 \div 298 = 0,395 \times 100 = 39,5\%$$

$$C = 3 \times 12 = 36 \div 298 = 0,120 \times 100 = 12\%$$

$$O = 9 \times 16 = \frac{144}{298} \div 298 = 0,483 \times 100 = 48,3\%$$

298 g/mol



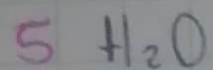
$$H = 2 \times 1 = 2 \div 98 = 0,020 \times 100 = 2\%$$

$$S = 1 \times 32 = 32 \div 98 = 0,326 \times 100 = 32,6\%$$

$$O = 4 \times 16 = \frac{64}{98} \div 98 = 0,653 \times 100 = 65,3\%$$

98 g/mol

cribe



$$\text{H} = 2 \times 1 = 2 \div 18 = 0,111 \times 100 = 11,1\%$$

$$\text{O} = 1 \times 16 = \frac{16}{18 \text{ g/mol}} \div 18 = 0,888 \times 100 = 88,8\%$$