Actividad modulo

1 calcula la masa molecular de las siguientes moleculas

a) $CaCO_3$

Ca = 40 g * 1 = 40 g
C = 12 g * 1 = 12 g
O = 16 g * 3 = 48 g +

100 = Tien $CaCO_3$

b) $Fe(NO_3)_3$

Fe = 56
N = 14
O = 16
56 + (3)(14) + (3)(3)(16))
56 + 42 + 144

242 = tiene $Fe(NO_3)_3$

c) HCl

Cl: 97.23
H: 2.76

36.46 = HCl

d) $Al(OH)_3$

Al = 27 g/mol
O = 16 g/mol x 3 = 48 g/mol
H = 1g/mol x 3 = 3 g/mol
27 + 16 + 3 =

78 = tiene $Al(OH)_3$

e) HNO_3

H = 1 g/mol
N = 14 g/mol
O = 16 g/mol

HNO_3 = 63

F) H_2SO_4

H = 1 g/mol
S = 32 g/mol
O = 16 g/mol

= 98 = H_2SO_4

g) $C_6H_{12}O_6$

S = 32 g/mol
O = 16 g/mol
H = 1 g/mol

98 = $C_6H_{12}O_6$

h) NaOH

C = 12 g/mol
H = 1 g/mol

$C_6H_{12}O_6 = 6 \times 12 + 12 \times 1 + 6 \times 16 =$
180

I) MgO

Mg = 24.3
O = 16

MgO = MO_3

j) $CuSO_4$

Cu = 64
O = 16
S = 32

160 uma de $CuSO_4$

k) NH_3

N = 14
H = 1(3) = 3
NH_3 = 17

l) C_6H_{14}

$C = 12 \times 6 = 72$
$H = 1 \times 14 = 14$

$C_6H_{14} = 86$

m) C_2H_2

$C = 12 \times 2 = 24$
$H = 1 \times 2 = 2$

$C_2H_2 = 26$

n) CO_2

$C = 12$ g/mol
$O_2 = 44$ g/mol

$CO_2 = 44.01$

o) Fe_2O_2

$Fe = (26) = 112$
$O = (16) = 32$
$= 144$ uma