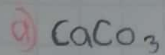


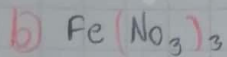
## Actividad modulo

1 calcula la masa molecular de las siguientes moléculas



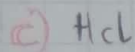
$$\begin{aligned}\text{Ca} &= 40 \text{ g} \times 1 = 40 \text{ g} \\ \text{C} &= 12 \text{ g} \times 1 = 12 \text{ g} \\ \text{O} &= 16 \text{ g} \times 3 = 48 \text{ g} +\end{aligned}$$

100 ~~g~~ tiene  $\text{CaCO}_3$



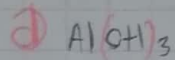
$$\begin{aligned}\text{Fe} &= 56 \\ \text{N} &= 14 \\ \text{O} &= 16 \\ 56 + (3)(14) + (9)(16)\end{aligned}$$

242 ~~g~~ tiene  $\text{Fe}(\text{NO}_3)_3$



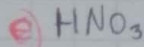
$$\begin{aligned}\text{Cl} &= 35.45 \\ \text{H} &= 1.008\end{aligned}$$

36.46 =  $\text{HCl}$



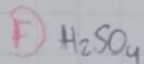
$$\begin{aligned}\text{Al} &= 27 \text{ g/mol} \\ \text{O} &= 16 \text{ g/mol} \times 3 = 48 \text{ g/mol} \\ \text{H} &= 1 \text{ g/mol} \times 3 = 3 \text{ g/mol} \\ 27 + 48 + 3 &= 78\end{aligned}$$

78 ~~g~~ tiene  $\text{Al}(\text{OH})_3$



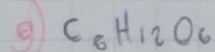
$$\begin{aligned}\text{H} &= 1 \text{ g/mol} \\ \text{N} &= 14 \text{ g/mol} \\ \text{O} &= 16 \text{ g/mol} \times 3 = 48\end{aligned}$$

$\text{HNO}_3 = 63$



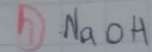
$$\begin{aligned}\text{H} &= 1 \text{ g/mol} \times 2 = 2 \\ \text{S} &= 32 \text{ g/mol} \\ \text{O} &= 16 \text{ g/mol} \times 4 = 64\end{aligned}$$

$2 + 32 + 64 = 98$



$$\begin{aligned}\text{C} &= 12 \text{ g/mol} \times 6 = 72 \\ \text{H} &= 1 \text{ g/mol} \times 12 = 12 \\ \text{O} &= 16 \text{ g/mol} \times 6 = 96\end{aligned}$$

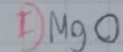
180 =  $\text{C}_6\text{H}_{12}\text{O}_6$



$$\begin{aligned}\text{Na} &= 23 \text{ g/mol} \\ \text{O} &= 16 \text{ g/mol} \\ \text{H} &= 1 \text{ g/mol}\end{aligned}$$

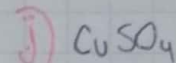
$$23 + 16 + 1 = 40$$

40 ~~g~~



$$\begin{aligned}\text{Mg} &= 24.3 \\ \text{O} &= 16\end{aligned}$$

$\text{MgO} = 40.3$

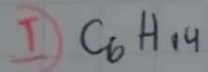


$$\begin{aligned}\text{Cu} &= 64 \\ \text{S} &= 32 \\ \text{O} &= 16 \times 4 = 64\end{aligned}$$

160 una da  $\text{CuSO}_4$



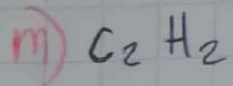
$$\begin{aligned}\text{N} &= 14 \\ \text{H} &= 1 \times 3 = 3 \\ \text{NH}_3 &= 17\end{aligned}$$



$$C = 12 \times 6 = 72$$

$$H = 1 \times 14 = 14$$

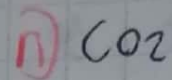
$$C_6H_{14} = 86$$



$$C = 12 \times 2 = 24$$

$$H = 1 \times 2 = 2$$

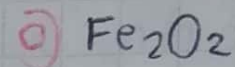
$$C_2H_2 = 26$$



$$CO = 12 \text{ g/mol}$$

$$O_2 = 44 \text{ g/mol}$$

$$CO_2 = 44.01$$



$$Fe = (26) = 112$$

$$O = (16) = 32$$

$$= 144 \text{ uma}$$