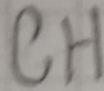


$$1. \text{ mol C} = \frac{92,3\text{g}}{12\text{g/mol}} = 7,691 \quad \text{mol } 7,691/7,7 = 1$$

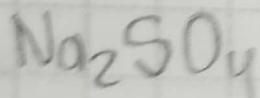
$$\text{mol H} = \frac{7,7\text{g}}{1\text{g/mol}} = 7,7 \quad \text{mol } 7,7/7,7 = 1$$



$$2. \text{ mol Na} = \frac{32,4\text{g}}{23\text{g/mol}} = 1,408 \quad \text{mol } 1,408/0,703 = 2$$

$$\text{mol S} = \frac{22,5\text{g}}{32\text{g/mol}} = 0,703 \quad \text{mol } 0,703/0,703 = 1$$

$$\text{mol O} = \frac{45,1\text{g}}{16\text{g/mol}} = 2,818 \quad \text{mol } 2,818/0,703 = 4$$



$$3. \text{ mol C} = \frac{48\text{g}}{12\text{g/mol}} = 4 \quad \text{mol } 4/0,4 = 1$$

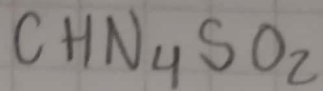
$$\text{mol H} = \frac{4\text{g}}{1\text{g/mol}} = 4 \quad \text{mol } 4/0,4 = 1$$

$$\text{mol N} = \frac{22,4\text{g}}{14\text{g/mol}} = 1,6 \quad \text{mol } 1,6/0,4 = 4$$

$$\text{mol S} = \frac{12,8\text{g}}{32\text{g/mol}} = 0,4 \quad \text{mol } 0,4/0,4 = 1$$

$$\text{mol O} = \frac{12,8 \text{ g}}{16 \text{ g/mol}} = 0,8$$

$$\text{mol } 0,8 / 0,4 = 2$$

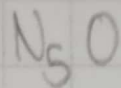


$$4. \text{ mol N} = \frac{0,079 \text{ g}}{14 \text{ g/mol}} = 5,642$$

$$\text{mol } 5,642 / 0,011 = 5$$

$$\text{mol O} = \frac{0,181 \text{ g}}{16 \text{ g/mol}} = 0,011$$

$$\text{mol } 0,011 / 0,011 = 1$$



$$5. \text{ mol Na} = \frac{21,6 \text{ g}}{23 \text{ g/mol}} = 0,939$$

$$\text{mol } 0,939 / 0,939 = 1$$

$$6. \text{ mol Cl} = \frac{33,3 \text{ g}}{35 \text{ g/mol}} = 0,951$$

$$\text{mol } 0,951 / 0,939 = 1$$

$$\text{mol O} = \frac{45,1 \text{ g}}{16 \text{ g/mol}} = 2,818$$

$$\text{mol } 2,818 / 0,939 = 3$$

