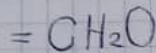


Formula Molecular

Ejercicios ②

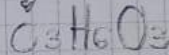
carbono	$\frac{40g}{12g/mol}$	3.3	$3.3/3.3 = 1$
hidrogeno	$\frac{6.7g}{1g/mol}$	6.7	$6.7/3.3 = 2$
oxigeno	$\frac{53.3}{16g/mol}$	3.331	$3.331/3.3 = 1$



masa = 90g/mol

C 1x12 = 12
H 1x2 = 2
O 1x16 = 16
= 30

$n = \frac{90}{30} = 3$



①

Masa = 1000g

Carbono 48% $\frac{48g}{12g/mol} = 4$

Hidrogeno 4% $\frac{4g}{1g/mol} = 4$

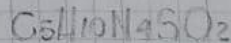
Nitrogeno 22.4% $\frac{22.4g}{14g/mol} = 1.6$

Azufre 12.8% $\frac{12.8g}{32g/mol} = 0.4$

oxigeno 12.8% $\frac{12.8g}{16g/mol} = 0.8$

C = $\frac{48}{12} = 4$
H = $\frac{4}{1} = 4$
N = $\frac{22.4}{14} = 1.6$
S = $\frac{12.8}{32} = 0.4$
O = $\frac{12.8}{16} = 0.8$

190



$n = \frac{1000}{190} = 5$

