

Formula empírica

Solución

1R/ C = 92,3%
H = 7,7%

Mol C $\frac{92,3 \text{ g}}{12 \text{ g/mol}}$ 7,691 $\frac{7,7}{7,7} = 1$

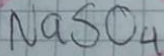
Mol H $\frac{7,7 \text{ g}}{1 \text{ g/mol}}$ 7,7 $\frac{7,7}{7,7} = 1$



2R/ Na $\frac{32,4 \text{ g}}{23 \text{ g/mol}}$ 1,408 $\frac{1,408}{0,703} = 2$

S $\frac{22,5 \text{ g}}{32 \text{ g/mol}}$ 0,703 $\frac{0,703}{0,703} = 1$

O $\frac{45,1 \text{ g}}{16 \text{ g/mol}}$ 2,818 $\frac{2,818}{0,703} = 4$



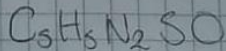
3R/ C $\frac{48 \text{ g}}{12 \text{ g/mol}}$ 4 $\frac{4}{0,8} = 5$

H $\frac{4 \text{ g}}{1 \text{ g/mol}}$ 4 $\frac{4}{0,8} = 5$

N $\frac{22,4 \text{ g}}{14 \text{ g/mol}}$ 1,6 $\frac{1,6}{0,8} = 2$

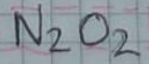
S $\frac{22,8 \text{ g}}{32 \text{ g/mol}}$ 0,7125 $\frac{0,7125}{0,8} = 0,89$

O $\frac{12,8 \text{ g}}{16 \text{ g/mol}}$ 0,8 $\frac{0,8}{0,8} = 1$



4R/ N $\frac{0,070 \text{ g}}{14 \text{ g/mol}}$ 0,005 $\frac{0,071}{0,005} = 14$

O $\frac{0,071 \text{ g}}{16 \text{ g/mol}}$ 0,0044 $\frac{0,071}{0,0044} = 16$



5R/ Na $\frac{2,939 \text{ g}}{23 \text{ g/mol}}$ 0,1278 $\frac{0,939}{0,1278} = 7$

Cl $\frac{33,3 \text{ g}}{35 \text{ g/mol}}$ 0,951 $\frac{0,981}{0,1278} = 7$

O $\frac{45,7 \text{ g}}{16 \text{ g/mol}}$ 2,856 $\frac{2,819}{0,1278} = 22$

