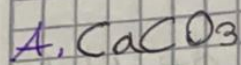


Taller
1. Calcula la masa molecular de las siguientes moléculas

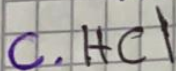
Solución



$$\text{Ca} = 40$$

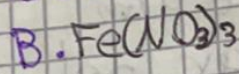
$$\text{C} = 12$$

$$\text{O}_3 = \frac{48}{100} \text{ uma}$$



$$\text{H} = 1$$

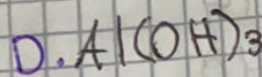
$$\text{Cl} = \frac{35}{36} \text{ uma}$$



$$\text{Fe} = 56$$

$$\text{N} = 42$$

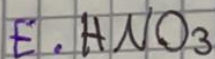
$$\text{O} = \frac{144}{242} \text{ uma}$$



$$\text{Al} = 27$$

$$\text{O} = 16$$

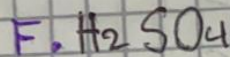
$$\text{H}_3 = \frac{3}{46} \text{ uma}$$



$$\text{H} = 1$$

$$\text{N} = 14$$

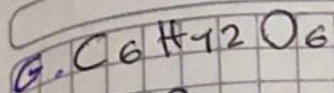
$$\text{O}_3 = \frac{48}{63} \text{ uma}$$



$$\text{H}_2 = 2$$

$$\text{S} = 32$$

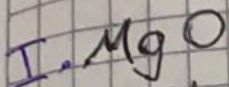
$$\text{O}_4 = \frac{64}{98} \text{ uma}$$



$$C_6 = 72$$

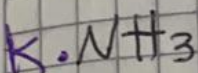
$$H_{12} = 12$$

$$O_6 = \frac{96}{180} \text{ uma}$$



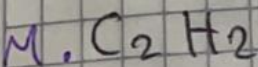
$$Mg = 24$$

$$O = \frac{16}{40} \text{ uma}$$



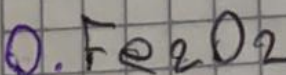
$$N = 14$$

$$H_3 = \frac{3}{17} \text{ uma}$$



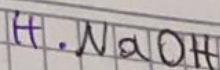
$$C_2 = 24$$

$$H_2 = \frac{2}{26} \text{ uma}$$



$$Fe_2 = 112$$

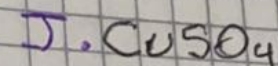
$$O_2 = \frac{32}{144} \text{ uma}$$



$$Na = 23$$

$$O = 16$$

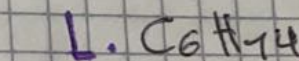
$$H = \frac{1}{40} \text{ uma}$$



$$Cu = 63$$

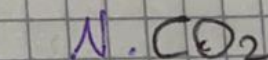
$$S = 32$$

$$O_4 = \frac{64}{159} \text{ uma}$$



$$C_6 = 72$$

$$H_{14} = \frac{14}{86} \text{ uma}$$



$$C = 12$$

$$O_2 = \frac{32}{44} \text{ uma}$$