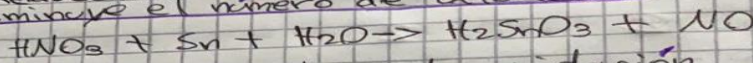


Balancedo por redox  
Número de oxidación

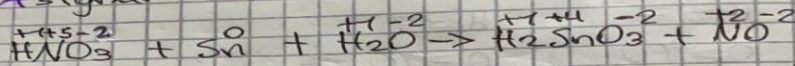
-2 -1 0 +1 +2 +3 +4 +5 +6 +7

Oxidación  
Aumenta el número de oxidación →

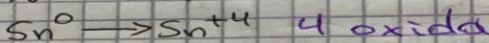
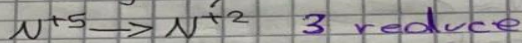
Reducción  
Disminuye el número de oxidación ←



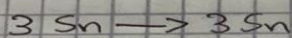
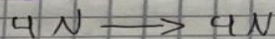
1. Asignar estados de oxidación



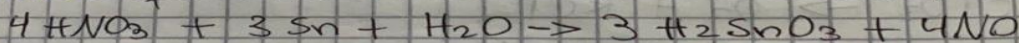
2. Identificar que elemento se oxida y cual se redujo y en cuanto



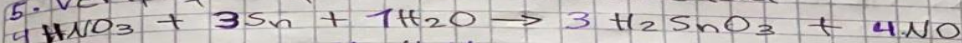
3. Cambie la diferencia entre estos



4. Ubique estos números en la ecuación



5. Verifique que este balanceada



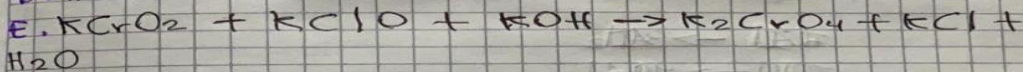
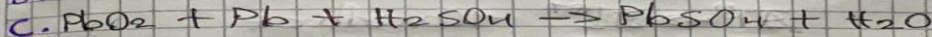
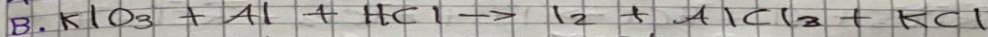
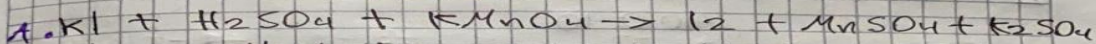
4 N 4

3 Sn 3

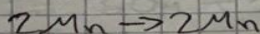
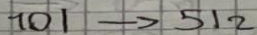
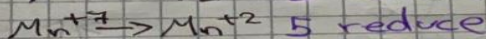
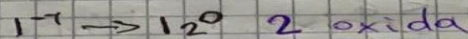
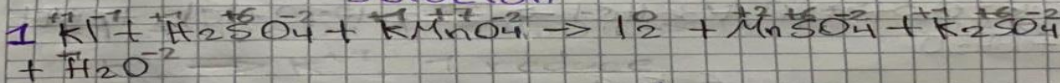
6 H 6

13 O 13

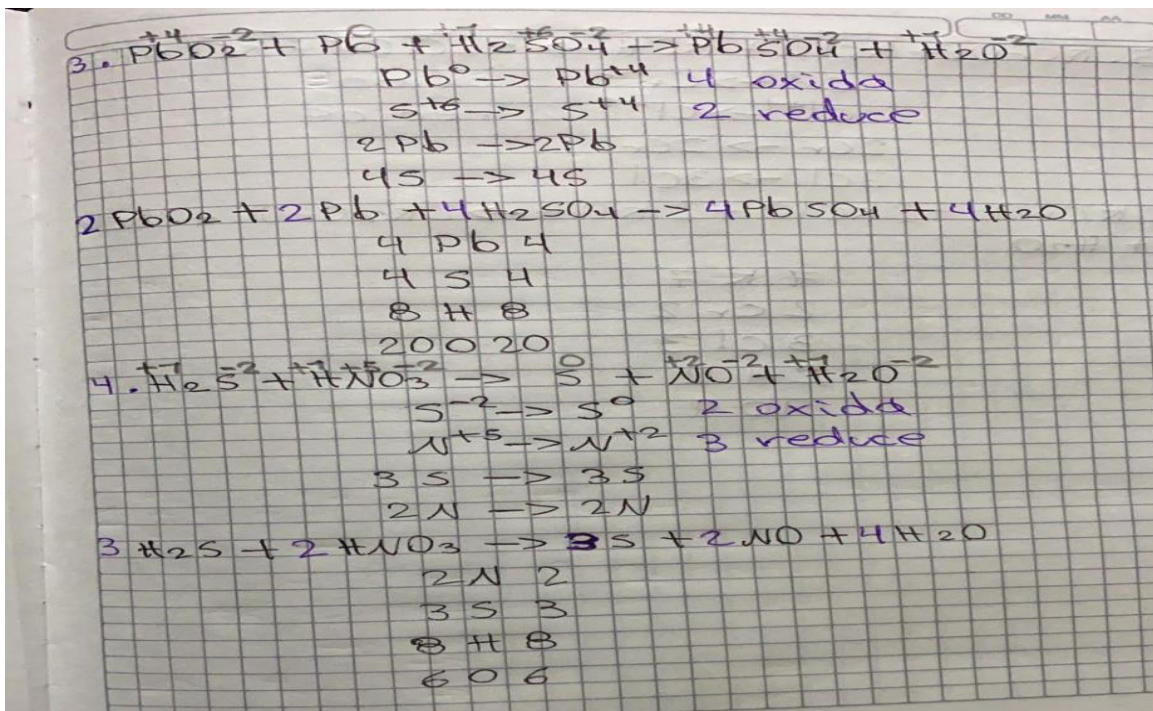
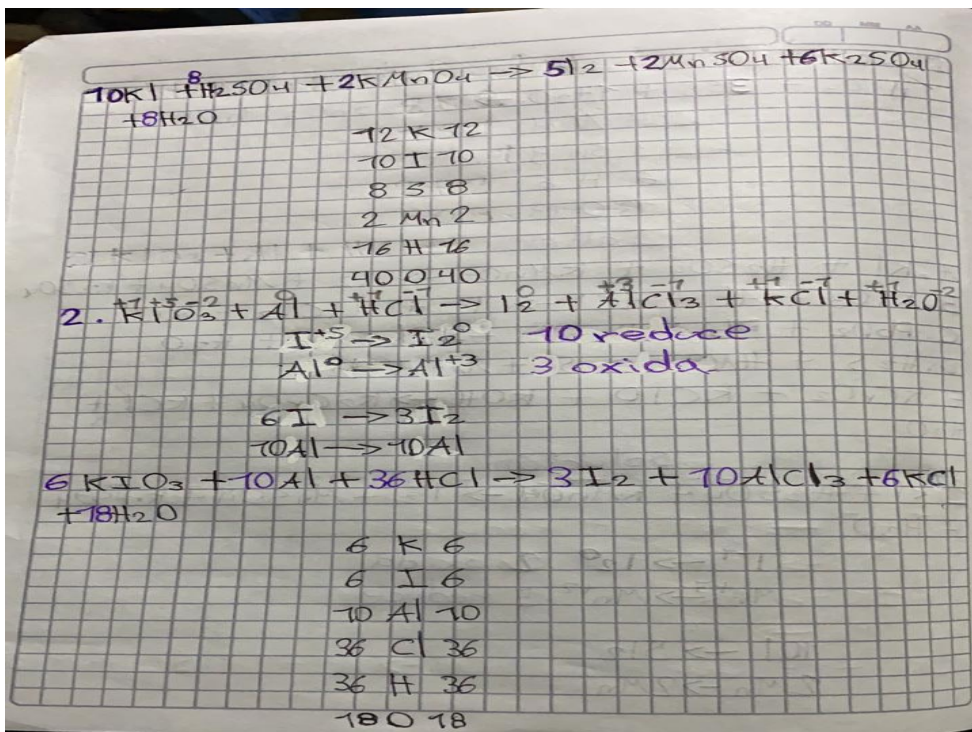
Taller

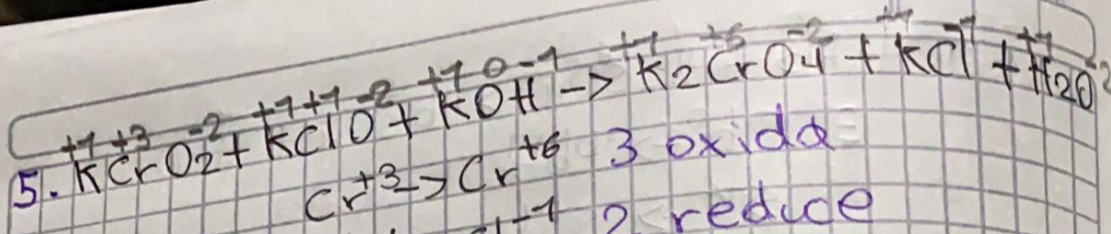


Solución

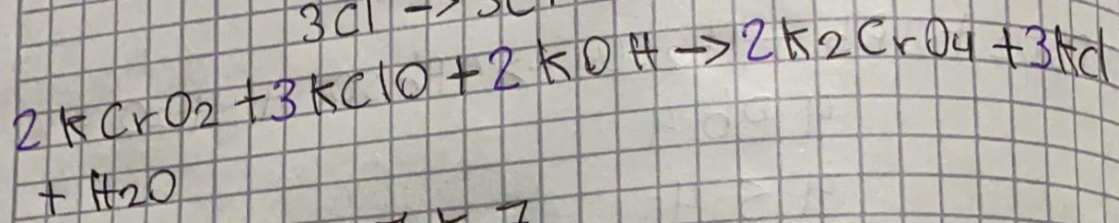
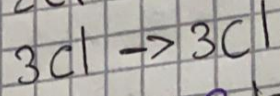
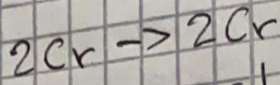








Cr<sup>+3</sup> → Cr<sup>+6</sup> 3 oxidize  
 Cl<sup>+1</sup> → Cl<sup>-1</sup> 2 reduce



7 K 7

2 Cr 2

3 Cl 3

2 H 2

9 O 9