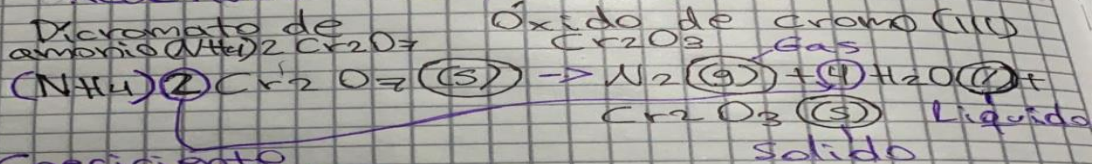


Comprender que es una ecuación y una reacción química

Balances de ecuaciones químicas

Representación simbólica o gráfica de una reacción química

Reactivos \rightarrow Productos



Coefficiente Estequiométrico

- S = Sólido
- G = Gaseoso
- L = Líquido
- Ac = Acuoso

- 1) $C_2H_6 + O_2 \rightarrow CO_2 + 3H_2O$

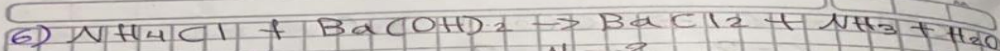
C = 2	C = 2
H = 6	H = 6
O = 2	O = 2
- 2) $Mg(s) + O_2(g) \rightarrow MgO(g)$

Mg = 1	Mg = 1
O = 1	O = 1
- 3) $P_4(s) + O_2(g) \rightarrow P_2O_5(g)$

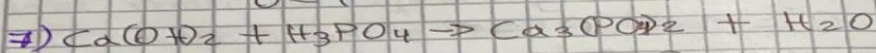
P = 4	P = 4
O = 6	O = 6
- 4) $C_3H_8(g) + O_2(g) \rightarrow CO_2(g) + H_2O(g)$

C = 3	C = 3
H = 8	H = 8
O = 6	O = 6
- 5) $HCl(ac) + Cr(s) \rightarrow CrCl_3(ac) + H_2(g)$

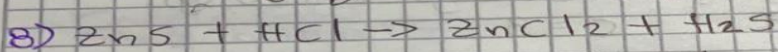
H = 6	H = 6
Cr = 2	Cr = 2
Cl = 6	Cl = 6



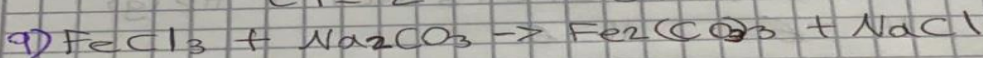
N = 2	N = 2
Ba = 2	Ba = 2
Cl = 2	Cl = 2
H = 6	H = 6
O = 2	O = 2



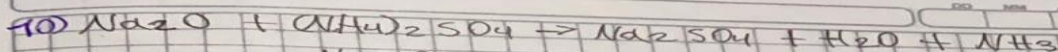
Ca = 3	Ca = 3
P = 6	P = 6
H = 6	H = 6
O = 6	O = 6



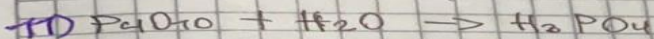
Zn = 1	Zn = 1	H = 2	H = 2
S = 1	S = 1		
Cl = 2	Cl = 2		



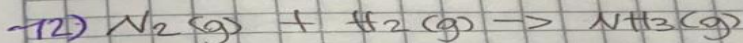
Fe = 2	Fe = 2
Cl = 6	Cl = 6
Na = 6	Na = 6
C = 3	C = 3
O = 9	O = 9



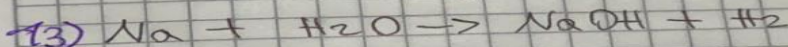
Na = 2	Na = 2
N = 2	N = 2
S = 1	S = 1
H = 6	H = 6
O = 4	O = 4



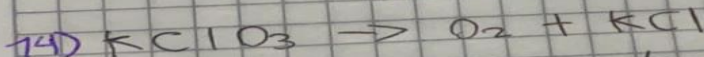
P = 4	P = 4
H = 2	H = 2
O = 10	O = 10



N = 2	N = 2
H = 5	H = 5



Na = 2	Na = 2
H = 4	H = 4
O = 2	O = 2



K = 2	K = 2
Cl = 2	Cl = 2
O = 5	O = 5