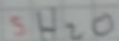
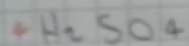
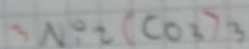
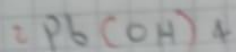
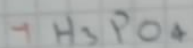


Actividad



Solución

$$\text{H} : 3 \times 1 = 3 \div 98 = 0.030 \times 100 = 3\%$$

$$\text{P} : 1 \times 31 = 31 \div 98 = 0.316 \times 100 = 31.6\%$$

$$\text{O} : 4 \times 16 = \frac{64}{98 \text{ g mol}} \div 98 = 0.653 \times 100 = \frac{65.3\%}{99.9}$$

$$\text{Pb: } 1 \times 207 = 207 \div 275 = 0.752 \times 100 = 75.2 \%$$

$$\text{O: } 4 \times 16 = 64 \div 275 = 0.232 \times 100 = 23.2 \%$$

$$\text{H: } 4 \times 1 = \frac{4}{275 \text{ g mol}} \div 275 = 0.014 \times 100 = \underline{1.4 \%}$$

99.8

$$\text{N: } 2 \times 58 = 116 \div 296 = 0.391 \times 100 = 39.1 \%$$

$$\text{C: } 3 \times 12 = 36 \div 296 = 0.121 \times 100 = 12.1 \%$$

$$\text{O: } 9 \times 16 = \frac{144}{296 \text{ g mol}} \div 296 = 0.486 \times 100 = \underline{48.6 \%}$$

99.8

$$\text{H: } 2 \times 1 = 2 \div 98 = 0.020 \times 100 = 2 \%$$

$$\text{S: } 1 \times 32 = 32 \div 98 = 0.326 \times 100 = 32.6 \%$$

$$\text{O: } 4 \times 16 = \frac{64}{98 \text{ g mol}} \div 98 = 0.653 \times 100 = \underline{65.3 \%}$$

99.9

$$\text{H: } 2 \times 1 = 2 \div 18 = 0.111 \times 100 = 11.1 \%$$

$$\text{O: } 1 \times 16 = \frac{16}{18 \text{ g mol}} \div 18 = 0.888 \times 100 = \underline{88.8 \%}$$

99.9