

cular de las siguientes moléculas.

a) CaCO_3

$$\begin{aligned}\text{Ca} &= 40 \cdot 1 = 40 \text{ uma} \\ \text{C} &= 12 \cdot 1 = 12 \text{ uma} \\ \text{O}_3 &= 16 \cdot 3 = 48 \text{ uma} \\ &= 100 \text{ uma}\end{aligned}$$

c) HCl

$$\begin{aligned}\text{H} &= 1 \cdot 1 = 1 \text{ uma} \\ \text{Cl} &= 35 \cdot 1 = 35 \text{ uma} \\ &= 36 \text{ uma}\end{aligned}$$

e) HNO_3

$$\begin{aligned}\text{H} &= 1 \cdot 1 = 1 \text{ uma} \\ \text{N} &= 14 \cdot 1 = 14 \text{ uma} \\ \text{O}_3 &= 16 \cdot 3 = 48 \text{ uma} \\ &= 63 \text{ uma}\end{aligned}$$

g) $\text{C}_6\text{H}_{12}\text{O}_6$

$$\begin{aligned}\text{C}_6 &= 12 \cdot 6 = 72 \text{ uma} \\ \text{H}_{12} &= 1 \cdot 12 = 12 \text{ uma} \\ \text{O}_6 &= 16 \cdot 6 = 96 \text{ uma} \\ &= 180 \text{ uma}\end{aligned}$$

i) MgO

$$\begin{aligned}\text{Mg} &= 24 \cdot 1 = 24 \text{ uma} \\ \text{O} &= 16 \cdot 1 = 16 \text{ uma} \\ &= 40 \text{ uma}\end{aligned}$$

b) $\text{Fe(NO}_3)_2$

$$\begin{aligned}\text{Fe} &= 56 \cdot 1 = 56 \text{ uma} \\ \text{N}_3 &= 14 \cdot 3 = 42 \text{ uma} \\ \text{O}_6 &= 16 \cdot 6 = 96 \text{ uma} \\ &= 194 \text{ uma}\end{aligned}$$

d) Al(OH)_3

$$\begin{aligned}\text{Al} &= 27 \cdot 1 = 27 \text{ uma} \\ \text{O}_3 &= 16 \cdot 3 = 48 \text{ uma} \\ \text{H}_3 &= 1 \cdot 3 = 3 \text{ uma} \\ &= 78 \text{ uma}\end{aligned}$$

f) H_2SO_4

$$\begin{aligned}\text{H}_2 &= 1 \cdot 2 = 2 \text{ uma} \\ \text{S} &= 32 \cdot 1 = 32 \text{ uma} \\ \text{O}_4 &= 16 \cdot 4 = 64 \text{ uma} \\ &= 98 \text{ uma}\end{aligned}$$

h) NaOH

$$\begin{aligned}\text{Na} &= 23 \cdot 1 = 23 \text{ uma} \\ \text{O} &= 16 \cdot 1 = 16 \text{ uma} \\ \text{H} &= 1 \cdot 1 = 1 \text{ uma} \\ &= 40 \text{ uma}\end{aligned}$$

j) CuSO_4

$$\begin{aligned}\text{Cu} &= 63 \cdot 1 = 63 \text{ uma} \\ \text{S} &= 32 \cdot 1 = 32 \text{ uma} \\ \text{O}_4 &= 16 \cdot 4 = 64 \text{ uma} \\ &= 159 \text{ uma}\end{aligned}$$

k) NH_3

$$\begin{aligned} \text{N} &= 14 \cdot 1 = 14 \text{ uma} \\ \text{H}_3 &= 1 \cdot 3 = 3 \text{ uma} \\ & \underline{17 \text{ uma}} \end{aligned}$$

m) C_2H_2

$$\begin{aligned} \text{C}_2 &= 12 \cdot 2 = 24 \text{ uma} \\ \text{H}_2 &= 1 \cdot 2 = 2 \text{ uma} \\ & \underline{26 \text{ uma}} \end{aligned}$$

o) Fe_2O_3

$$\begin{aligned} \text{Fe}_2 &= 56 \cdot 2 = 112 \text{ uma} \\ \text{O}_3 &= 16 \cdot 3 = 48 \text{ uma} \\ & \underline{160 \text{ uma}} \end{aligned}$$

l) C_6H_{14}

$$\begin{aligned} \text{C}_6 &= 12 \cdot 6 = 72 \text{ uma} \\ \text{H}_{14} &= 1 \cdot 14 = 14 \text{ uma} \\ & \underline{86 \text{ uma}} \end{aligned}$$

n) CO_2

$$\begin{aligned} \text{C} &= 12 \cdot 1 = 12 \text{ uma} \\ \text{O}_2 &= 16 \cdot 2 = 32 \text{ uma} \\ & \underline{44 \text{ uma}} \end{aligned}$$