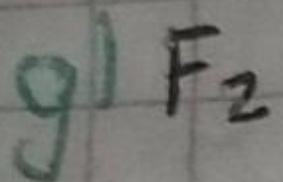
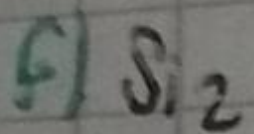
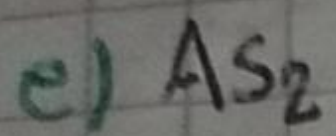
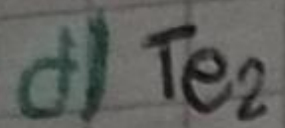
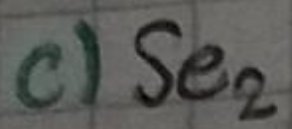
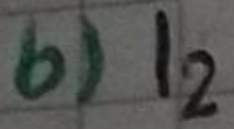
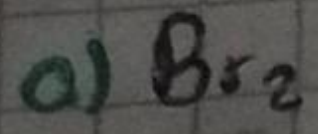
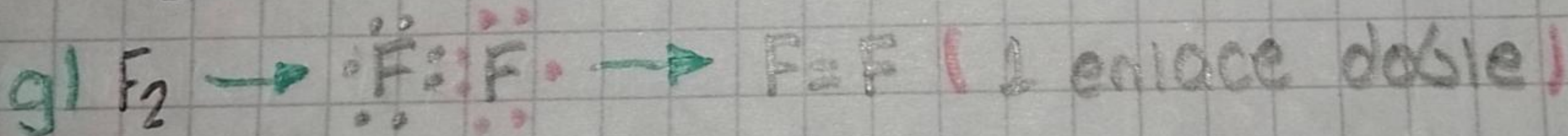
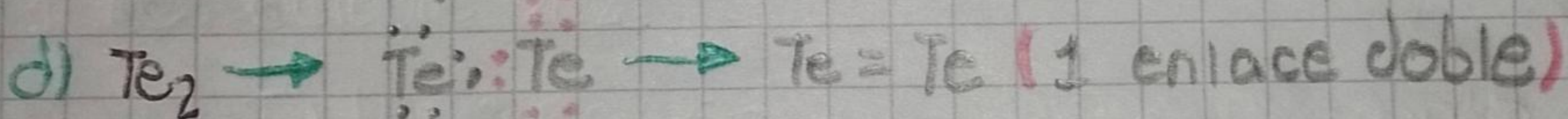
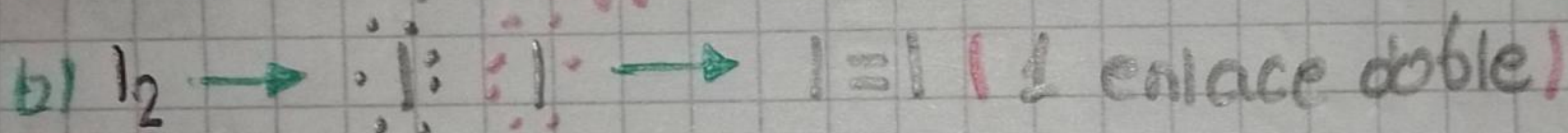


Ejercicios:



Solución:

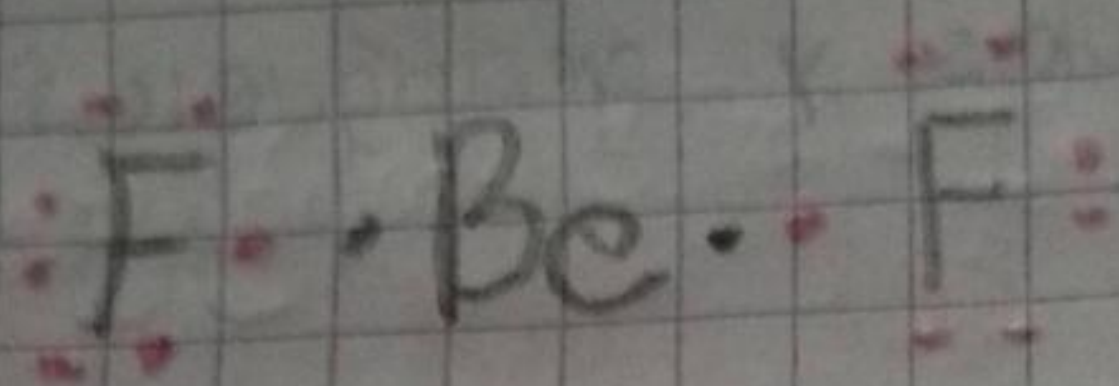


• Todos son no polares

SOLUCIÓN

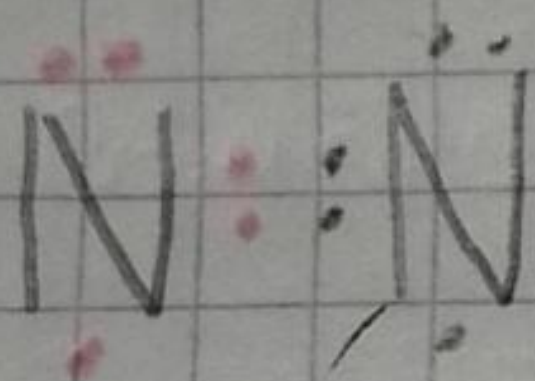
1.

a) BeF_2



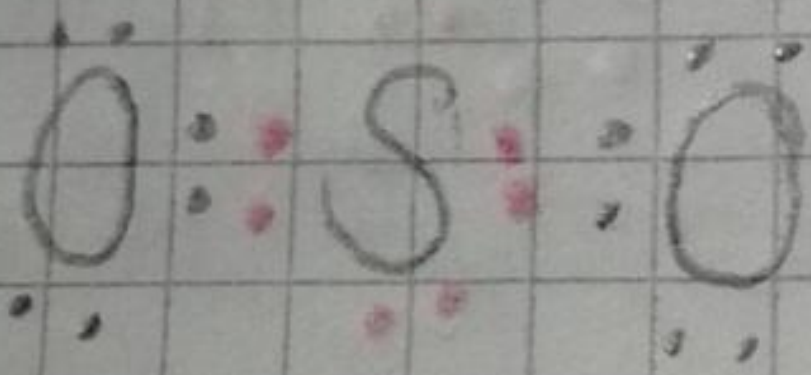
Tipo de enlace: Covalente polar = $3.98 - 1.57 = 2.41$

b) N_2



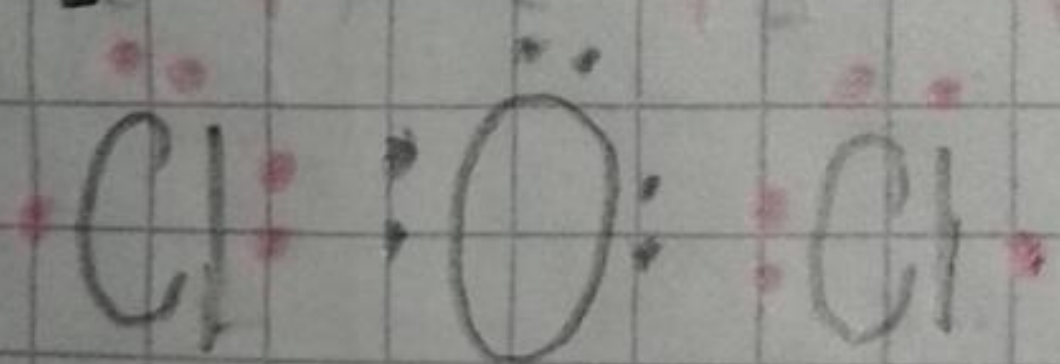
Tipo de enlace: Covalente no polar = $3.04 - 3.04 = 0$

c) SO_2



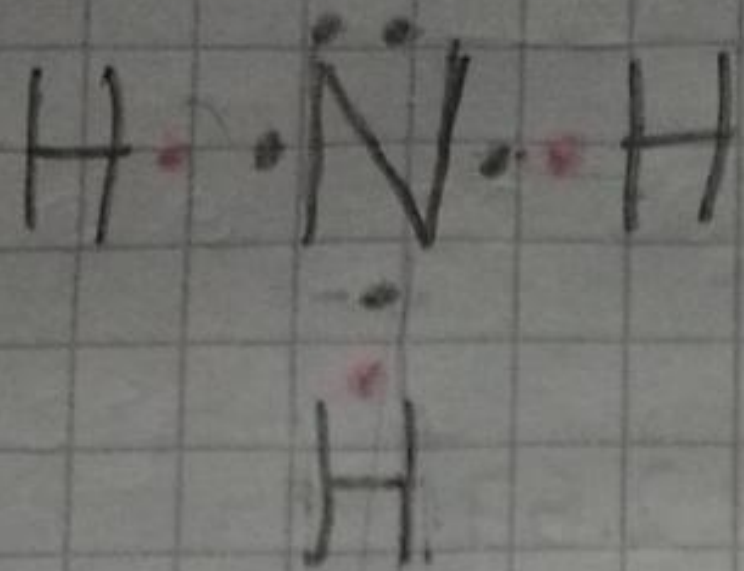
Tipo de enlace: Covalente polar = $3.44 - 2.58 = 0.86$

d) Cl_2O



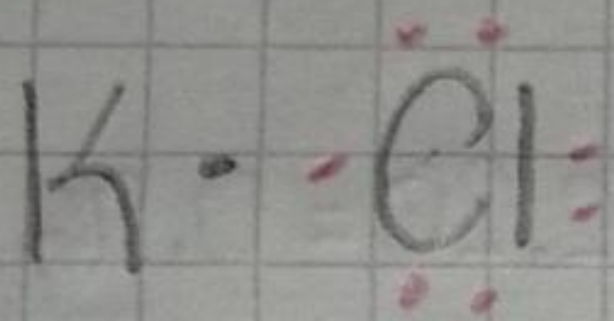
Tipo de enlace: Covalente no polar = $3.44 - 3.16 = 0.28$

e) NH_3



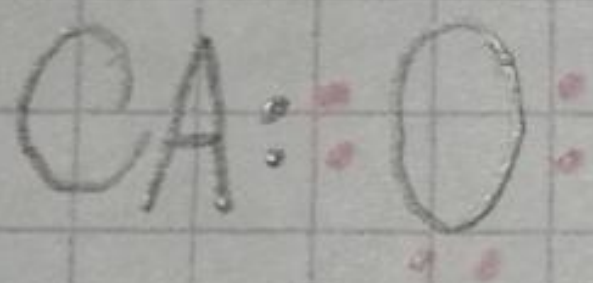
Tipo de enlace: Covalente polar = $3.04 - 2.20 = 0.84$

f) KCl



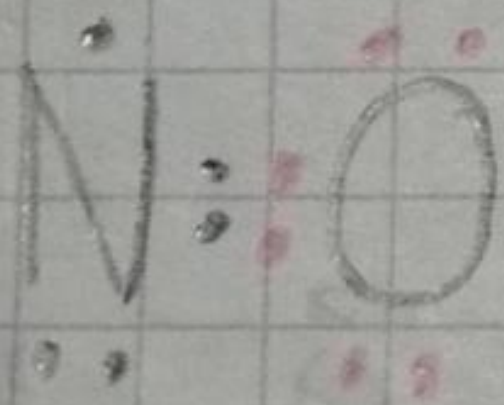
Tipo de enlace: Iónico = $3.16 - 0.82 = 2.34$

g) CaO



Tipo de enlace: Iónico = $3.44 - 1.00 = 2.44$

h) NO



Tipo de enlace: Covalente no polar = $3.44 - 3.04 = 0.4$

