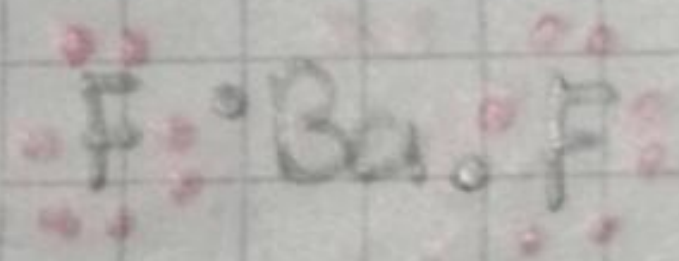
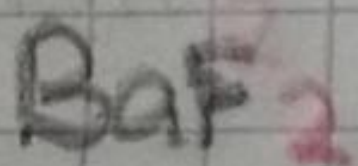


## TAREA

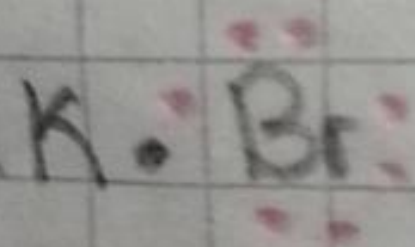
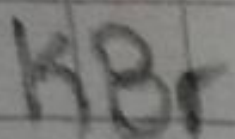
■ Graficar los siguientes compuestos con la estructura de Lewis y electronegatividades.

$BaF_2$	Ba (metal)	F (no metal)
$KBr$	K (metal)	Br (no metal)
$FeCl_3$	Fe (metal)	Cl (no metal)
$AlI_3$	Al (metal)	I (no metal)
$CaBr_2$	Ca (metal)	Br (no metal)

## SOLUCIÓN

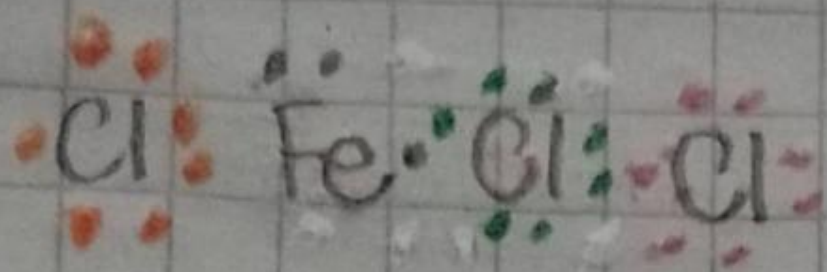
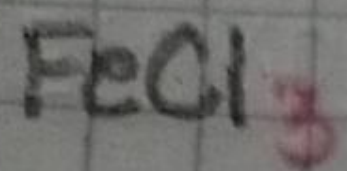


$$BaF_2 = 3.98 - 0.89 = 3.09$$

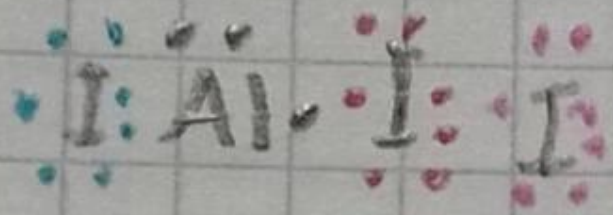
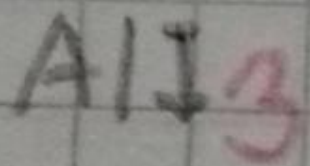


$$KBr = 2.96 - 0.82 = 2.14$$

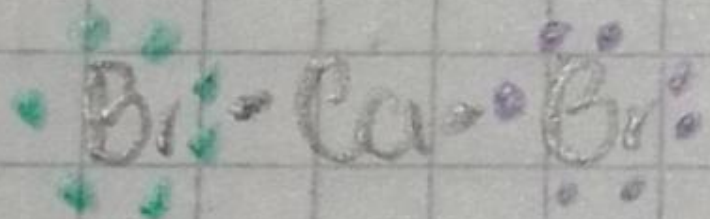
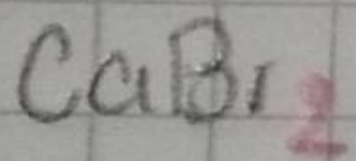




$$\text{FeCl}_3 = 3.16 - 1.83 = 1.33$$



$$\text{AlI}_3 = 2.66 - 1.61 = 1.05$$



$$\text{CaBr}_2 = 2.96 - 1.00 = 1.96$$