

$$1. \text{ mol C} = \frac{48 \text{ g}}{12 \text{ g/mol}} = 4$$

$$\text{mol } 4/0,4 = 10$$

$$\text{mol H} = \frac{4 \text{ g}}{1 \text{ g/mol}} = 4$$

$$\text{mol } 4/0,4 = 10$$

$$\text{mol N} = \frac{22,4 \text{ g}}{14 \text{ g/mol}} = 1,6$$

$$\text{mol } 1,6/0,4 = 4$$

$$\text{mol S} = \frac{12,8 \text{ g}}{32 \text{ g/mol}} = 0,4$$

$$\text{mol } 0,4/0,4 = 1$$

$$\text{mol O} = \frac{12,8 \text{ g}}{16 \text{ g/mol}} = 0,8$$

$$\text{mol } 0,8/0,4 = 2$$

fórmula empírica: $C_4H_4N_4S_1O_2$
massa molecular: $C_4H_4N_4S_1O_2$

$$C = 12 \times 4 = 48$$

$$H = 1 \times 4 = 4$$

$$N = 14 \times 4 = 56$$

$$S = 32 \times 1 = 32$$

$$O = 16 \times 2 = 32$$

$$250$$

$$C = 4 \times 12 = 48$$

$$H = 4 \times 1 = 4$$

$$N = 4 \times 14 = 56$$

$$S = 4 \times 8 = 32$$

$$O = 4 \times 2 = 8$$

$$100\% = 4$$

$$250$$

$$2. \text{ mol C} = \frac{40 \text{ g}}{12 \text{ g/mol}} = 3,3$$

$$\text{mol } 3,3/3,3 = 1$$

$$\text{mol H} = \frac{6,7 \text{ g}}{1 \text{ g/mol}} = 6,7$$

$$\text{mol } 6,7/3,3 = 2$$

$$\text{mol O} = \frac{53,3 \text{ g}}{16 \text{ g/mol}} = 3,3$$

$$\text{mol } 3,3/3,3 = 1$$

fórmula empírica: CH_2O
massa molecular: CH_2O

$$C = 3 \times 1 = 3$$

$$H = 3 \times 2 = 6$$

$$O = 3 \times 1 = 3$$

$$C = 12 \times 1 = 12$$

$$H = 1 \times 2 = 2$$

$$O = 16 \times 1 = 16$$

30

$$\frac{009}{30} = 3$$

30