

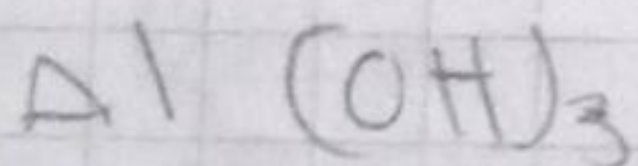
NORMALIDAD

$$N = \frac{\text{equivalente gr Sólido}}{\text{Litro de solución}}$$

solución

• Obtenga la normalidad de $\text{Al}(\text{OH})_3$ con 450 ml y 9.50 gramos

• $\text{Al}(\text{OH})_3$ $m = 9,50$
 $v = 450 \text{ ml} / 1000 = 0,45 \text{ L}$



$$\text{PM} = \text{H} = 3 \times 1 = 3$$

$$\text{O} = 3 \times 16 = 48$$

$$\text{Al} = 1 \times 27 = 27$$

$$\text{PM} = 78 \text{ g/mol}$$

$$m = 9,50 \text{ gramos} \quad / \quad N = \frac{(9,50) (3)}{(78) (0,45)}$$

$$\text{PM} = 78 \text{ g/mol}$$

$$v = 0,45 \text{ L} \quad / \quad N = \frac{28,5}{35,1} = 0,81$$

$$N = 0,81$$