

SOLUCIÓN

Date / /

① C = 92,3 Mol C = $\frac{92,3g}{12g/mol} = 7,6916$

H = 7,7 Mol H = $\frac{7,7}{1g/mol} = 7,7$

Mol = 7,6916 = 59,13

Mol = 7,7 = 7,6916 59 = **CH**

② Na = 32,4 Mol Na = $\frac{32,4g}{23g/mol} = 1,4086$

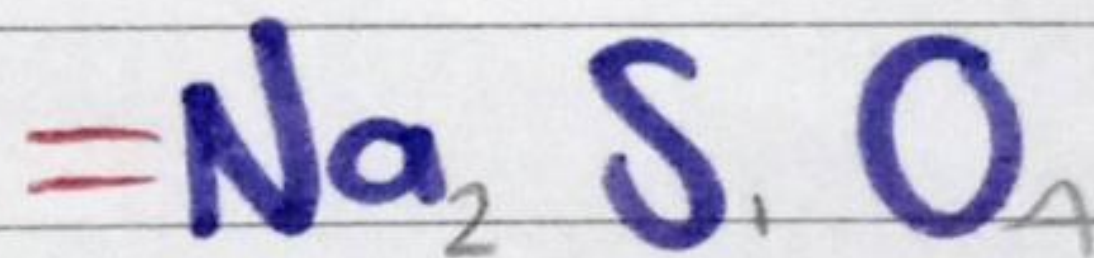
S = 22,5 Mol S = $\frac{22,5g}{32g/mol} = 0,69$

O = 45,1 Mol O = $\frac{45,1g}{16g/mol} = 2,818$

- 1,4086 = 2

- 0,6968 $\times 0,6968 = 1$

- 2,818 = 4



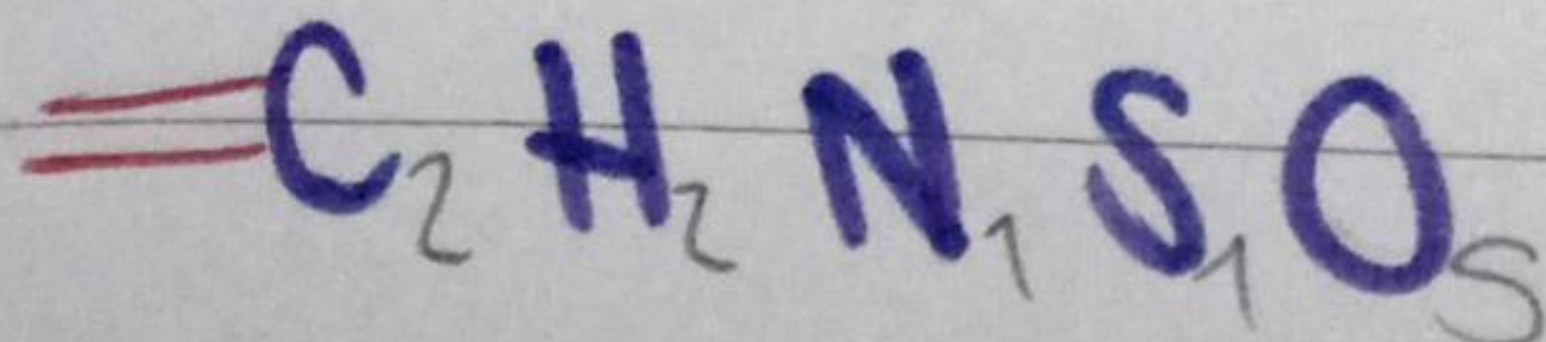
③ H = 4 Mol H = $\frac{4g}{1g/mol} = 4$ = 16

N = 22,4 Mol N = $\frac{22,4}{14g/mol} = 1,6$ = 0,64

S = 12,8 Mol S = $\frac{12,8g}{32g/mol} = 0,4$ $\times 0,4 = 0,16$

O = 12 Mol O = $\frac{12,8g}{16g/mol} = 12,8$ = 5,12

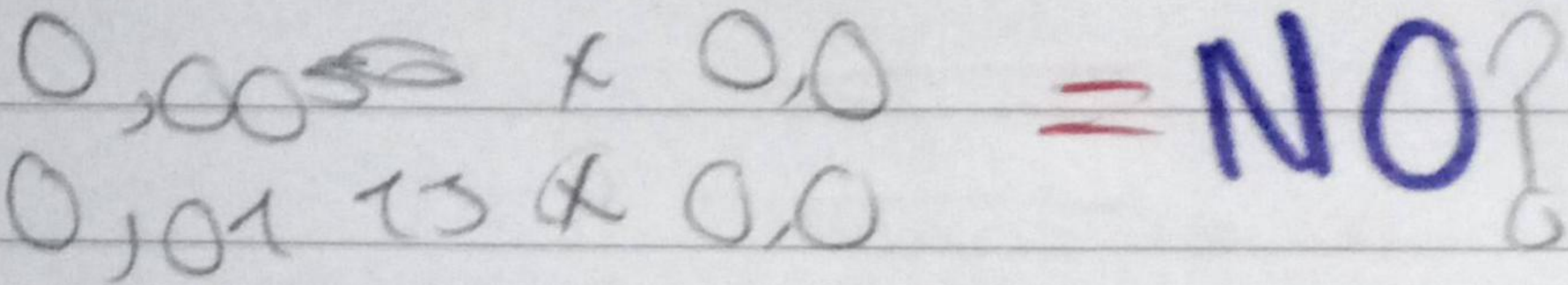
C = 48 Mol C = $\frac{48g}{12g/mol} = 4$ = 1,6



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$$\textcircled{iv} \text{ N} = 0 \text{ ?} = \frac{0,0399 \text{ g}}{14 \text{ g/mol}} = 0,00285$$

$$\text{O} = 0 \text{ ?} = \frac{0,181 \text{ g}}{16 \text{ g/mol}} = 0,0113$$



$$\begin{array}{l} \textcircled{v} \text{ S } 21,6 \text{ Mol } \frac{21,6 \text{ g}}{32 \text{ g/mol}} = 0,675 \quad / \quad 0,6 = 0,4 \\ \text{Cl } 33,3 \text{ Mol } \frac{33,3 \text{ g}}{35 \text{ g/mol}} = 0,9514 \quad / \quad \times 0,6 = 0,6 \\ \text{O } 4,9 \text{ Mol } \frac{4,9 \text{ g}}{16 \text{ g/mol}} = 0,306 \quad / \quad 0,6 = 1,9 \end{array}$$

