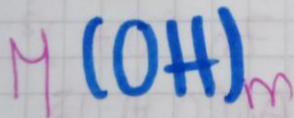


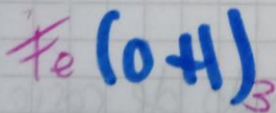
Comprender de manera adecuada los Hidroxidos

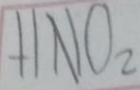
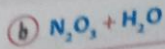
Formulación

Se escribe a la izquierda el metal (M) que es el más electropositivo y a la derecha el grupo hidroxilo (OH) y después se intercambian los números de oxidación



Ejemplo Hidroxido de Hierro III

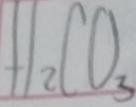
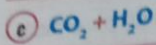




Acido Nitroso

Oxido de Nitrogeno (III)

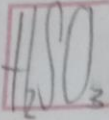
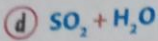
dioxonitrato de Nitrogeno



Acido carbonico

Acido trioxocarbono (IV)

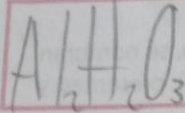
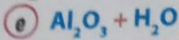
trioxocarbonato de hidrogeno



Acido Sulfuroso

Acido trioxidosulfuro

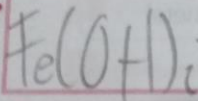
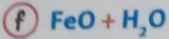
trioxidosulfato de hidrogeno



Oxido de Aluminio

Oxido de aluminio (III)

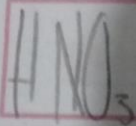
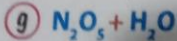
hidroxido de Aluminio



Hidroxido ferroso

Hidroxido de Hierro (II)

dioxido de Hierro



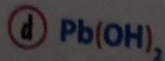
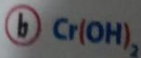
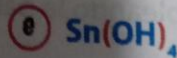
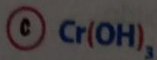
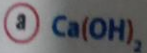
Acido Nitrico

Acido trioxonitrico

trioxonitrato de hidrogeno

En tu cuaderno

2 Nombra las siguientes sustancias con las tres nomenclaturas.



3 Escribir la fórmula de las sustancias con sus estados de oxidación.

a) Hidróxido níquelico

b) Hidróxido aúrico

e) Dihidróxido de Cinc