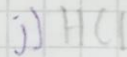
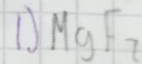
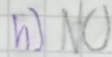
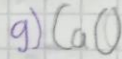
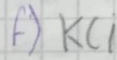
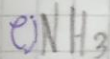
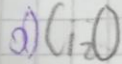
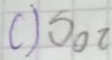
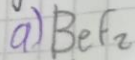


Taller.

1. Indique el tipo de enlace que presenta cada una de las siguientes moléculas:



2. De las siguientes moléculas: F_2 , CO_2 , C_2H_4 , C_2H_2 , H_2O , C_6H_6 , NH_3

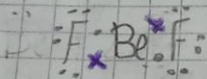
a) ¿Cuáles tienen todos los enlaces sencillos o simples?

b) ¿Dónde existe algún doble enlace?

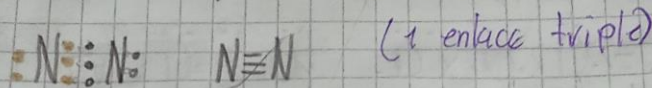
c) ¿Dónde existe algún triple enlace?

Solución

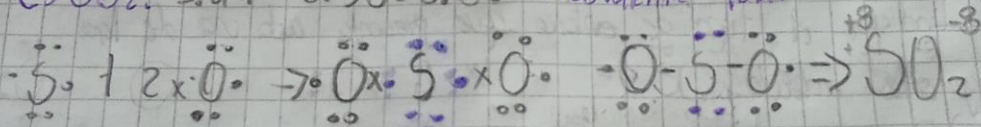
1 a) $\text{BeF}_2: 3.98 - 1.57 = 2.41 = \text{Enlace iónico}$



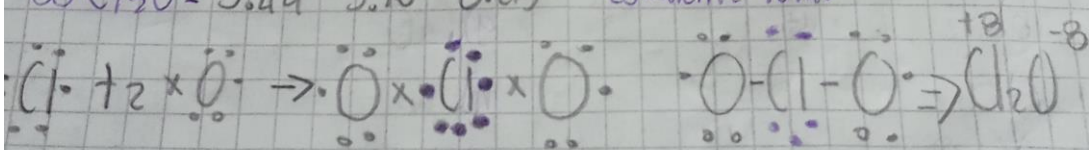
b) $\text{N}_2: 3.04 - 3.04 = 0 = \text{Covalente no polar}$



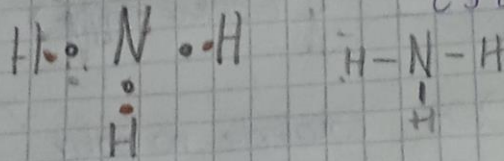
c) $\text{SO}_2: 3.44 - 2.58 = 0.86 = \text{Covalente polar}$



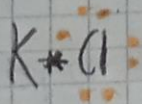
d) $\text{Cl}_2\text{O}: 3.44 - 3.16 = 0.28 = \text{Covalente polar}$



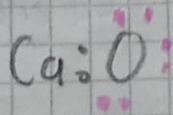
e) $\text{NH}_3: 2.20 - 3.04 = -0.84 = \text{Covalente no polar}$
 (3 enlaces simples)



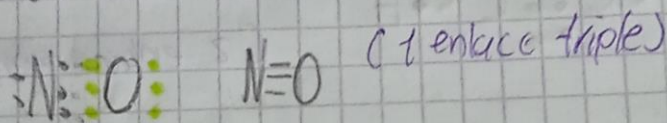
f) $\text{KCl}: 3.16 - 0.82 = 2.34 = \text{Enlace iónico}$



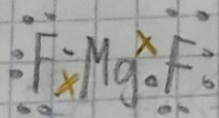
g) $\text{CaO}: 3.44 - 1.00 = 2.44 = \text{Enlace iónico}$



h) $\text{NO}: 3.44 - 3.01 = 0.43 = \text{Covalente no polar}$



i) $\text{MgF}_2: 3.98 - 1.31 = 2.67 = \text{enlace iónico}$



j) $\text{HCl}: 3.16 - 2.20 = 0.96 = \text{covalente polar}$

