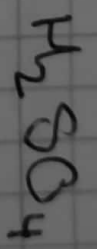


$$M = 49.034 \text{ g/mol}$$



$$V = 250 \text{ ml} \approx 1000 = 0.25 \text{ L}$$

$$H \times 2 \times 1 = 2$$

$$S \times 1 \times 32 = 32$$

$$O \times 4 \times 16 = 64$$

$$PM = 112$$

$$M = \frac{0.437}{28} = 0.015$$

## Unidades químicas de concentración

Molaridad

$$M = \frac{\text{Moles de soluto}}{\text{Litros de solución}}$$

$$M = \frac{m}{(PM)(V)}$$

Actividad

1  $M =$

$m = 82.5 \text{ g}$

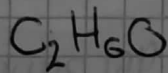
$V = 0.45 \text{ L}$

$C \times 12 = 24$

$H \times 1 = 16 +$

$O \times 1 = 16 = 6$

$PM = 46 \text{ g/mol}$



$$M = \frac{82.5}{(46)(0.45)} = \frac{82.5}{20.7} = M = 4$$

$M = 0.086$

2  $m = 4.78 \text{ g/mol}$

$V = 7 \text{ L}$

$$M = \frac{4.78}{7} = M = 0.682$$

3  $m = 3.65 \text{ g/mol}$



$V = 2 \text{ L}$

$H = 1 + 1 = 2$

$Cl = 35.5 \times 1 = 35.5$

$PM = 36.5$

$$M = \frac{3.65}{73} = 0.05 \text{ M}$$