

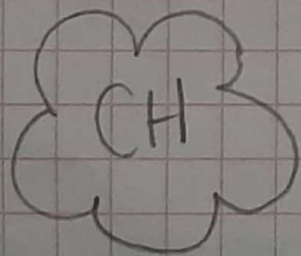
# EJERCICIOS:

$$\textcircled{1} \text{ C} = 92,3\%$$

$$\text{H} = 7,7\%$$

$$\text{Mol C} = \frac{92,3\text{g}}{12\text{g/mol}} = 7,691 / 7,7 = 1 \quad \text{aproximarlo}$$

$$\text{Mol H} = \frac{7,7\text{g}}{1\text{g/mol}} = 7,7 / 7,7 = 1$$



Emmis

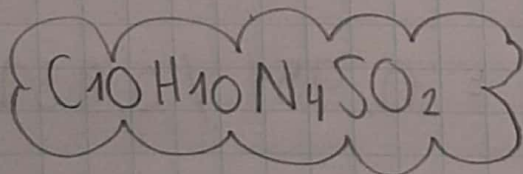
$$\text{Mol C} = \frac{48\text{g}}{12\text{g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol H} = \frac{4\text{g}}{1\text{g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol N} = \frac{22,4\text{g}}{14\text{g/mol}} = 1,6 / 0,4 = 4$$

$$\text{Mol S} = \frac{12,8\text{g}}{32\text{g/mol}} = 0,4 / 0,4 = 1$$

$$\text{Mol O} = \frac{12,8\text{g}}{16\text{g/mol}} = 0,8 / 0,4 = 2$$



$$\textcircled{4} \text{ N} = 0,079$$

$$\text{O} = 0,181$$



② Na = 32,4 %

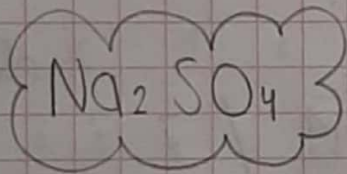
S = 22,5 %

O = 45,1 %

$$\text{Mol Na} \quad \frac{32,4\text{g}}{23\text{g/mol}} = 1,408 / 0,703 = 2$$

$$\text{Mol S} \quad \frac{22,5\text{g}}{32\text{g/mol}} = 0,703 / 0,703 = 1$$

$$\text{Mol O} \quad \frac{45,1\text{g}}{16\text{g/mol}} = 2,818 / 0,703 = 4$$



③ C = 48 %

H = 4 %

N = 22,4 %

S = 12,8 %

O = 12,8 %

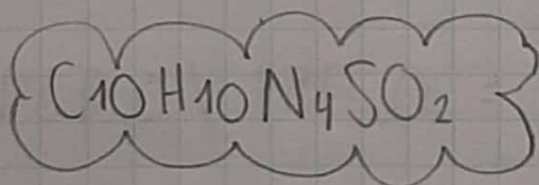
$$\text{Mol C} = \frac{48\text{g}}{12\text{g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol H} = \frac{4\text{g}}{1\text{g/mol}} = 4 / 0,4 = 10$$

$$\text{Mol N} = \frac{22,4\text{g}}{14\text{g/mol}} = 1,6 / 0,4 = 4$$

$$\text{Mol S} = \frac{12,8\text{g}}{32\text{g/mol}} = 0,4 / 0,4 = 1$$

$$\text{Mol O} = \frac{12,8\text{g}}{16\text{g/mol}} = 0,8 / 0,4 = 2$$



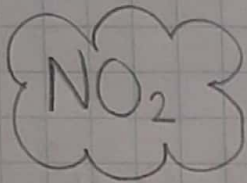
$$\textcircled{4} \quad \text{N} = 0,079$$

$$\text{O} = 0,181$$



$$\text{Mol N} = \frac{0,079\text{g}}{14\text{g/mol}} = 0,005 / 0,005 = 1$$

$$\text{Mol O} = \frac{0,181\text{g}}{16\text{g/mol}} = 0,011 / 0,005 = 2$$



- ⑤ Na = 21,6 %  
Cl = 33,3 %  
O = 45,1 %

$$\text{Mol Na} = \frac{21,6\text{g}}{23\text{g/mol}} = 0,939 / 0,939 = 1$$

$$\text{Mol Cl} = \frac{33,3\text{g}}{35\text{g/mol}} = 0,951 / 0,939 = 1$$

$$\text{Mol O} = \frac{45,1\text{g}}{16\text{g/mol}} = 2,818 / 0,939 = 3$$

