

EJERCICIOS..

① H_3PO_4 = Ácido fosfórico

$$H = 3 \times 1 = 3 \div 98 = 0.030 \times 100 = 3\%$$

$$P = 1 \times 31 = 31 \div 98 = 0.316 \times 100 = 31.6\%$$

$$O = 4 \times 16 = \frac{64}{98} \div 98 = 0.653 \times 100 = \frac{65.3\%}{99.9\%}$$

② $Pb(OH)_4$ = Hidróxido de Plomo (IV)

$$Pb = 1 \times 207 = 207 \div 275 = 0.752 \times 100 = 75.2\%$$

$$O = 4 \times 16 = 64 \div 275 = 0.232 \times 100 = 23.2\%$$

$$H = 4 \times 1 = \frac{4}{275} \div 275 = 0.014 \times 100 = \frac{1.4\%}{99.8\%}$$

③ $Ni_2(CO_3)_3$ = Carbonato de Niquel (III)

$$Ni = 2 \times 58 = 116 \div 296 = 0.391 \times 100 = 39.1\%$$

$$C = 3 \times 12 = 36 \div 296 = 0.121 \times 100 = 12.1\%$$

$$O = 9 \times 16 = \frac{144}{296} \div 296 = 0.486 \times 100 = \frac{48.6\%}{99.8\%}$$

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④ H_2SO_4 = Ácido sulfúrico

$$H = 2 \times 1 = 2 \div 98 = 0.020 \times 100 = 2.0\%$$

$$S = 1 \times 32 = 32 \div 98 = 0.326 \times 100 = 32.6\%$$

$$O = 4 \times 16 = \frac{64}{98} \div 98 = 0.653 \times 100 = \frac{65.3\%}{99.9}$$

⑤ H_2O = Agua

$$H = 2 \times 1 = 2 \div 18 = 0.111 \times 100 = 11.1\%$$

$$O = 1 \times 16 = \frac{16}{18} \div 18 = 0.888 \times 100 = \frac{88.8\%}{99.9}$$

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